

Research Article

Structural Insight into Chitosan Supports Functionalized with Nanoparticles

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The incorporation of suspensions of nanoparticles functionalized with gallic acid (GA) was used as a strategy to obtain nanocomposite active films with different both chitosan : tripolyphosphate (CH : TPP) and nanoparticles:chitosan (N : CH) ratios. The thermal analysis carried out by modulated differential scanning calorimetry (MDSC) allowed observing the shift of an endothermic event towards higher temperatures with a greater N : CH ratio. Analyzing ATR-FTIR spectra through principal component analysis (PCA) can be inferred that the incorporation of the nanoparticles produced a discrimination of the samples into clusters when the region $1400\text{--}1700\text{ cm}^{-1}$ was considered. The decrease in crystalline size with the inclusion of nanoparticles (N_A and N_B) proved the existence of interactions among CH, TPP, and GA, resulting in a more amorphous structure. The positron annihilation lifetime spectroscopy (PALS) technique was adequate to correlate the glass transition temperatures (T_g) obtained by using the MDSC technique with parameters τ_2 and I_2 ascribed to the annihilation of positrons in the interface. The cross section of nanocomposites obtained by scanning electron microscopy (SEM) clearly showed a homogeneous distribution of the nanoparticles without aggregation, suggesting their compatibility with the CH matrix. By virtue of the obtained results, the nanocomposites with the greatest nanoparticle proportion and the highest TPP concentration attained significant modifications in relation to CH matrices because of the crosslinking of the biopolymer with GA and TPP.

1. Introduction

Using the nanotechnology, new forms of polymeric supports with incorporation of active loaded nanoparticles can be developed to modulate their properties. The chitosan (CH) is a cationic biopolymer obtained for deacetylation of chitin [1]. The sodium tripolyphosphate (TPP) is a multivalent anion capability of interacting with the amino groups of the CH and forming an ionic network interconnected by means of the ionotropic gelation technique, allowing the attainment of nanoparticles [2]. These structures facilitate the nanoencapsulation of active compounds. The gallic acid (3,4,5-trihydroxybenzoic acid) is a phenolic acid that can be found in several fruits and in processed food and beverages [3]. This active compound has recognized biological properties

as an anti-inflammatory, antimutagenic, antioxidant, and antimicrobial agent [4–6]. The possibility of incorporating phenolic compounds into antimicrobial packaging proved to be of wide interest in such applications [7]. Since these components have antimicrobial activity, their release into food matrices may have a significant impact on shelf-life extension.

Considering this background, we have recently developed chitosan-based films with the incorporation of functionalized nanoparticles [2]. These nanocomposite systems demonstrated to have antimicrobial properties, and the ability to modulate the delivery of GA, hence, could be potential candidates to be used for packaging in food preservation.

According to dos Santos et al. [8], from the microstructural point of view, the incorporation of nanoparticles produces an interruption of the chitosan chains causing

a decrease in the free volume of the system. The polymer matrix, in the vicinity of the nanoparticles, generates interphases with properties different from the bulk. Consequently, the properties of the polymer such as free volume and chain conformation experienced deviations from the bulk behavior [9].

The versatility of the positron annihilation lifetime spectroscopy technique (PALS) for the study of polymers allows an integrated approach in conjunction with complementary techniques such as ATR-FTIR, MDSC, and SEM to establish composition-structure relationships at a molecular level [10].

The incorporation of suspensions of nanoparticles functionalized with GA was used as a strategy to obtain nanocomposite active films with different both chitosan: tripolyphosphate (CH:TPP) and nanoparticles:chitosan (N:CH) ratios.

The influence of the CH:TPP (v/v) and N:CH (v/v) ratios on the molecular mobility and the thermal and microstructural film properties was analyzed by means of ATR-FTIR, XRD, PALS, and SEM techniques. In addition, thermal behavior of the nanocomposites was studied by using differential scanning calorimetry (MDSC), and the observed variation in the thermal transition process was discussed in relation to the PALS parameters.

2. Materials and Methods

2.1. Materials. Chitosan (CH) from crab shells, with a deacetylation degree of 85% and a MW of 4.8×10^4 Da, was purchased from Polymar Ciência e Nutrição (Fortaleza, Brazil). Sodium tripolyphosphate (TPP) and acetic acid were supplied by Anedra (Buenos Aires, Argentina), and gallic acid (GA) was provided by Sigma-Aldrich Chemical Company (St. Louis, MO, USA).

2.2. Preparation of Nanoparticles. Chitosan 0.75% (w/w) was solubilized in 0.75% (v/v) acetic acid solution at 20°C under continuous agitation. Then, the corresponding amounts of GA were added to the solutions until reaching the concentration of $50 \text{ mg}_{\text{GA}}/\text{g}_{\text{CH}}$ (expressed as mg per gram CH). Sodium tripolyphosphate, used as a crosslinking agent, was prepared at the concentration of 2% (w/v) [11]. Nanoparticles were prepared by the ionotropic gelation technique based on a modified method described by Calvo et al. [12]. TPP solution was added dropwise under continuous stirring into the solution containing CH and GA until reaching CH:TPP ratios of 5 and 3. Afterward, the obtained suspensions were homogenized at 13,500 rpm for 10 min by using an Ultraturrax T-25 Janke & Kunkel (IKA-Labortechnik, Germany).

2.3. Characterization of Nanoparticle Suspensions. The zeta potential (ZP) and particle hydrodynamic size of the nanoparticles were determined by a Zetasizer Nano-ZS Malvern Instruments (Worcestershire, England) equipped with a digital correlator Model ZEN3600. The morphological characteristics of the nanoparticles were analyzed by using a transmission electron microscope (JEM 1200EX II Jeol, Japan) equipped with a digital camera (ES500W Erlangshen CCD). Prior to

examination, the nanoparticles were isolated by centrifugation (Beckman Coulter Optima L-100XP Floor Centrifugation System) at 20,000 rpm for 20 min. The supernatants were discarded, and the particles were washed several times with distilled water, redispersed, and diluted with distilled water.

2.4. Percentage Encapsulation Efficiency (PEE). To determine the percentage encapsulation efficiency, the synthesized nanoparticles were separated from the suspension by centrifugation at 20,000 rpm for 20 min (Beckman Coulter Optima L-100XP Floor Centrifugation System, California, USA). Samples were filtered using $0.45 \mu\text{m}$ sterile nylon (Millipore, Bedford, MA, USA), and the concentration of GA was quantified by using a Spectrophotometer DU650 (Beckman, USA). The concentration of the active agent in both the supernatant and the precipitate was calculated according to Lamarra et al. [2].

2.5. Preparation of Nanocomposite Matrices. A chitosan solution 1.5% (w/w) was prepared by solubilization in 1.5% (v/v) acetic acid solution as described previously [2].

Different proportions of suspension of nanoparticles (N_A and N_B) were dispersed in the CH solutions to obtain relations N:CH 1 : 2, 1 : 1, and 2 : 1 (v/v). The nanocomposites were prepared by casting and drying at 37°C until reaching a constant weight.

From now on, the formulations with suspensions N_A and N_B will be named as CH: N_A and CH: N_B , indicating in each case the N:CH ratio: 1 : 2, 1 : 1, and 2 : 1.

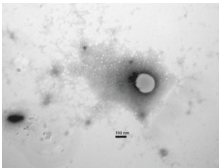
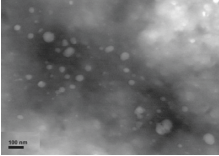
2.6. Thermal Analysis. Thermal properties of nanocomposites and CH films were conducted by using a DSC model Q100 controlled by a TA 5000 module (TA Instruments, New Castle, Delaware, USA), with a quench cooling accessory, under a N_2 atmosphere (20 ml/min). Samples were analyzed between -50°C and 200°C at a heating rate of $10^\circ\text{C}/\text{min}$ with a modulation period of 60 s and temperature amplitude of 0.5°C . After the first scan was completed, the sample was cooled until -50°C and then a second scan was recorded. All results were the average of two replicates. The total, reversing and nonreversing signals were determined. From the thermograms, the following information was obtained: T_p , the peak temperature ($^\circ\text{C}$); ΔH , the enthalpy ($\text{J}\cdot\text{g}^{-1}$, dry basis); and T_g , the glass transition temperature ($^\circ\text{C}$), by means of Universal Analysis V1.7F software (TA Instruments, New Castle, USA).

2.7. Structural Studies. The examination of the morphology of the nanocomposites was performed with low vacuum by means of a FEI model Quanta 200 scanning electron microscope (The Netherlands).

Before the analyses, nanocomposite films were cryogenically frozen in liquid nitrogen, fractured and mounted on bronze stubs through a double sided adhesive carbon tape.

Specimens were examined at an acceleration voltage of 15 kV without any metal or carbon coating.

TABLE 1: Zeta potential (ZP, mV), nanoparticle diameters (nm), TEM micrographs, polydispersity index (PDI), and percentage encapsulation efficiency (PEE) of N_A and N_B suspensions. Variables were chitosan concentration (w/w), chitosan:tripolyphosphate ratio (CH:TPP), and gallic acid concentration expressed as $\text{mg}_{\text{GA}}/\text{g}_{\text{CH}}$.

Suspension	CH (%)	CH:TPP	GA ($\text{mg}/\text{g}_{\text{CH}}$)	ZP (mV)*	Diameter (nm)	TEM	PDI	PEE
N_A	0.75	5	50	46.6 (3.2) ^b	453.1 (30.0) ^b		0.61 (0.02) ^b	80.9 (2.1) ^a
N_B	0.75	3	50	27.7 (1.2) ^a	140.3 (7.1) ^a		0.18 (0.01) ^a	80.4 (1.8) ^a

*Different letters in the same column indicate significant differences ($P < 0.05$) between samples.

The structure of the nanocomposite films was also evaluated by means of X-ray diffraction in a Philips PW3710, X'Pert Pro P Analytical ModelPW3040/60 (Almelo, The Netherlands) at room temperature. The CuK radiation (1.542 \AA) was generated at 30 mA and 40 kV, recording the relative intensity in the scattering range of (2θ) $3\text{--}60^\circ$ with a step size 2θ 0.02° .

The study of the contribution of each component to the interactions was carried out by means of ATR-FTIR technique. Spectra were recorded by a Nicolet, iS10 Thermo Scientific (Madison, USA), by the accumulation of 32 scans at 4 cm^{-1} resolution in the wavenumber range $4000\text{--}400 \text{ cm}^{-1}$. Specimens were placed onto the diamond ATR crystal using a top-plate and pressure-arm accessories (Smart iTX accessory) for the Nicolet™ iS™ 10 (Thermo Scientific™, Madison, USA). Data were analyzed by using the software Omnic 8 (Thermo Scientific, Madison, USA). The spectral deconvolution of the peaks within the region $1700\text{--}1500 \text{ cm}^{-1}$ (amide I and amide II bands) was performed by using curve-fitting algorithms. Inverted second derivative spectra were used to estimate the number, position, and relative contribution of individual element components of amide I and amide II. The software iterated the curve-fitting process by adjusting the high and width of the peaks to achieve the best Gaussian-shaped curves that fit the original spectrum.

2.8. Positrons Annihilation Lifetime Spectroscopy. Positron annihilation lifetime spectroscopy has been the most informative annihilation spectroscopy in characterizing the free volume of materials, especially thin films. Using a low-energy positron beam, free volume size and free volume concentration can be measured from the Ps lifetime and intensity data [13, 14]. Usually, the third lifetime component (τ_3) is associated with the means radius of the free volume sites, and the intensity I_3 is related to the concentration of the free volume sites.

Samples in the form of disks of sufficient thickness (1.3–1.5 mm) to ensure complete stopping of the positrons were placed on either side of the Kapton foils embedding the ^{22}Na source. The lifetime spectra were recorded in a fast-fast coincidence PALS setup with a time resolution (full-width-at-half-maximum, FWHM) of 260 ps [10]. Each sample was

measured three times with individual lifetime spectra having 3×10^6 integral counts. From the analysis, two discrete lifetimes and a continuous one corresponding to the long-lived component were obtained. This longest lifetime gives the ortho-positronium (o-Ps) lifetime component represented by the continuous distribution $1/\tau_3$. Positron lifetime spectra were also analyzed using the POSITRONFIT program, giving three discrete lifetime components (τ_1 , τ_2 , and τ_3 , resp.).

2.9. Statistical Analysis. For statistical analysis, Systat-software (SYSTAT, Inc., Evanston, IL, USA) version 10.0 was applied. Analysis of variance (ANOVA), linear regressions, and Fisher LSD mean comparison test were used. The significance levels used were 0.05.

Multivariate analysis and data preprocessing were conducted on the FTIR-ATR spectra, using The Unscrambler® software (version 8, CAMO, Norway). In order to evaluate the spectral differences among samples containing different proportions of nanoparticles, a principal component analysis (PCA) was performed over ATR-FTIR spectra covering the range $1700\text{--}1400 \text{ cm}^{-1}$. Taking into account the spectral differences associated with the presence of nanoparticles (N_A and N_B) in the PCA scores plot, a PLS model was calibrated to determine the relation between the TPP concentration and the absorbance corresponding to 900 cm^{-1} .

3. Results and Discussion

3.1. Characterization of Nanoparticles. According to Lamarra et al. [2], chitosan is a compound which combines the electrostatic stabilization due to its positive charge and the steric stabilization because of its polymeric nature. In this way, CH nanoparticles were obtained as a result of molecular linkages formed by inducing the ionotropic gelation between chitosan amino group and TPP phosphates. The concentration of CH, gallic acid ($\text{mg}_{\text{GA}}/\text{g}_{\text{CH}}$), and CH:TPP ratio used to obtain the suspensions (N_A and N_B), as well as zeta potential, size values, polydispersity index (PDI), and TEM micrographs of the formed nanoparticles are shown in Table 1. As it can be seen, the lower the CH:TPP ratio, the smaller

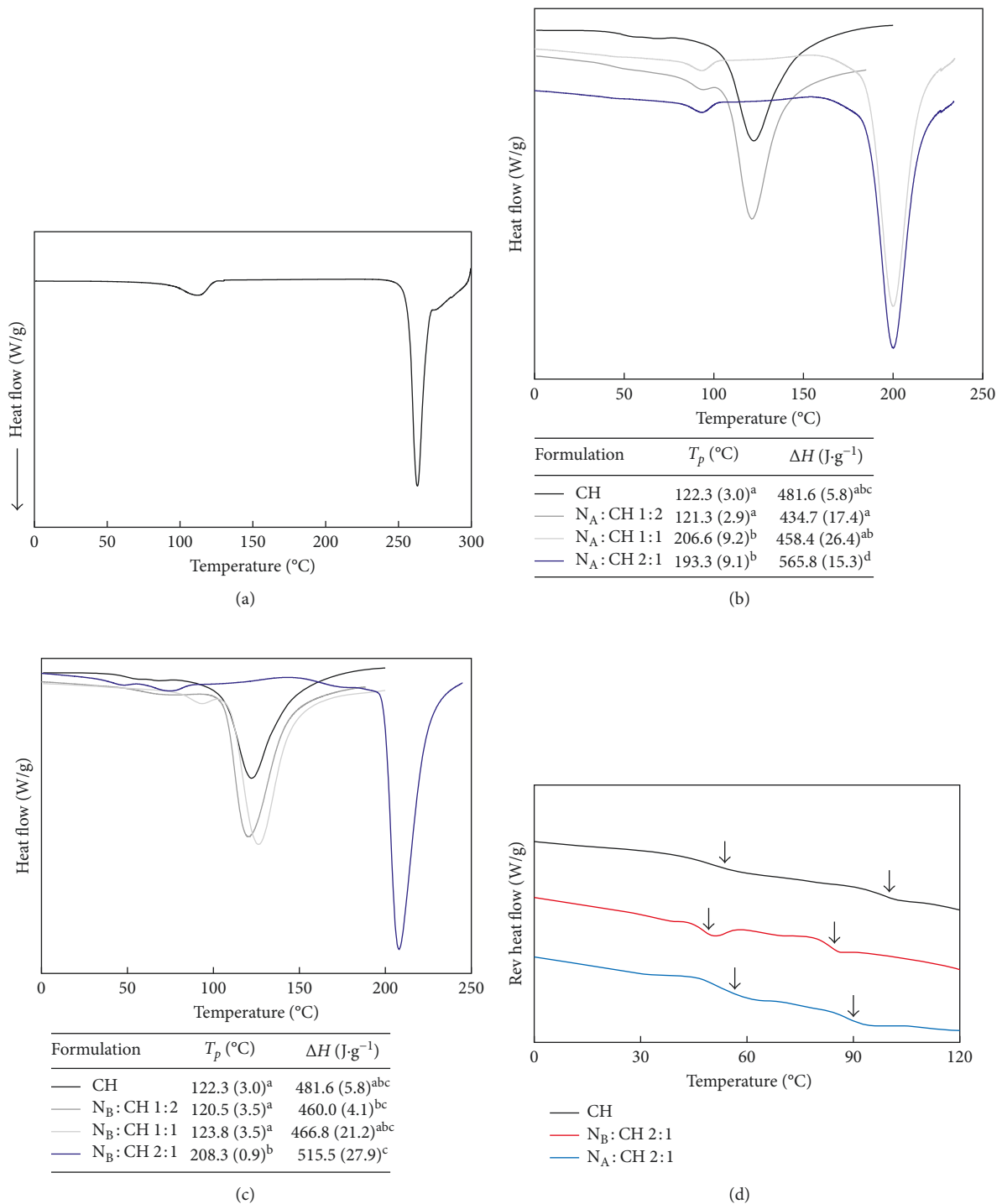


FIGURE 1: MDSC thermograms of gallic acid powder (a) and nanocomposite matrices with different N : CH ratios: N_A (CH : TPP 5) (b), and N_B (CH : TPP 3) (c). Table shows the parameters associated with the endothermic event. Reversible heat flow curves show the glass transition temperature of CH (black full line), N_A : CH 2 : 1 (blue full line), and N_B : CH 2 : 1 (red full line) (d). The arrows indicate the inflection points corresponding to the T_{gs} .

the particle diameter and the lower the ZP obtained. The morphological analysis examined by TEM confirmed the nanosize of the formulations and indicated that the nanoparticles were roughly spherical in shape. Likewise, N_B suspension exhibited a PDI value lower than 0.2, indicating a narrow particle size distribution (PDI < 0.5)

Meanwhile, PDI of N_A suspension was associated with the higher diameter obtained for these nanoparticles, giving a value higher than 0.5.

3.2. *Thermal Analysis.* DSC curves for GA displayed two endothermic transitions at 112.1 and 262.8°C (Figure 1(a)).

The first transition corresponds to the water loss and the second is related to melting of the mixture of isomers of GA, considering that according to da Rosa et al. [15], the melting point of these occurs at the following temperatures: 100°C, 221°C, and 250°C.

The curves obtained for the nanocomposites showed an endothermic event associated with the removal of water present in the matrix. A shift of the endothermic peak toward higher temperatures was observed as a consequence of the incorporation of higher proportion of nanoparticles (Figures 1(b) and 1(c)). The higher the CH:TPP ratio and, hence, the lower the crosslinking capacity, the greater the shift of the peak was. The location of the peak at the highest temperature could be attributed to the gallic acid free in the matrix whose melting is produced at 250°C [15, 16].

Therefore, variations in the enthalpy values evidenced by DSC curves and the position of the peaks related to the water loss may reflect the molecular and physical changes caused by the interaction of the GA molecules with chitosan polymer (table in Figure 1).

According to Shavi et al. [17], the absence of an independent melting peak of the active compound in the thermogram indicates that the drug is dispersed uniformly at a molecular level in the particles. In the nanocomposite matrices coexist two phases, an enriched phase with nanoparticles loaded with GA and a phase composed of chitosan and free GA.

For formulations $N_A:CH$ 1 : 2, $N_B:CH$ 1 : 2, and $N_B:CH$ 1 : 1, the endothermic transition appear at 125°C, approximately indicating the formation of an inclusion complex. Similar results were reported by several authors [15, 18].

As the N:CH ratio increases, the amount of free GA was higher, resulting in a shift of the endothermic transition to higher temperatures, this effect being more pronounced in formulations containing N_A nanoparticles. This result could be attributed to the fact that N_B contains a higher concentration of TPP that promotes ionic interactions between negative charges of crosslinker and positively charged groups.

Sharma et al. [19] informed that the glass transition of the nanocomposite matrices is used to observe the effect of the nanofillers on the thermal properties. The modifications in thermal properties as a function of the loading are also used to infer the changes in the molecular packing of polymer chains as a result of interfacial interaction between polymer and nanofillers [19].

Figure 1(d) shows the glass transition temperature (T_g) obtained from MDSC thermograms of the reversible heat flow of the nanocomposites. The T_g of chitosan film was observed at 91°C on average, which is in agreement with that informed by other authors [20, 21].

3.3. Film Microstructure. The micrographs of the CH film obtained by SEM (not shown) exposed a smooth and compact matrix, which was modified by the addition of nanoparticles. On the other hand, the cross section of nanocomposites clearly showed a homogeneous distribution

of the nanoparticles without aggregation, suggesting compatibility with the CH matrix (Figure 2).

Micrographs of the nanocomposites exhibited a uniform dispersion of the nanoparticles forming a discontinuous phase acting as a filler immersed in a CH continuous network irrespective of the N:CH ratio (Figure 2). In addition, the micrographs confirmed the minor size of the N_B in relation to N_A .

According to Ghozali and Haryono [22] with the inclusion of smaller particles, it can enhance the distribution of the nanoparticles in the polymer matrix.

3.4. DRX. As can be seen in Figure 3, CH matrix exhibited peaks localized at $2\theta = 8.2, 11.2, 18.1,$ and 23.0° which are characteristics of the polymer. The inclusion of nanoparticles (N_A and N_B) induced modifications in the reflection intensities of CH peaks centered at 11.2° and at 8.2° , indicating changes in the structure. It is recognized that the width of X-ray diffraction peak is related to the size of the crystallite, and the appearance of broadened peaks is the result from imperfect crystal Sujima-Anbu et al. [23]. The crystalline domains' size of CH and nanocomposites matrices were calculated by using the Debye-Scherrer equation [24]. The diffraction peak selected to calculate the crystalline size was the reflection near 18.1° [2]. The results obtained are summarized in the table in Figure 3, which exhibits a CH crystalline domain size of 20.9 nm. The decrease in crystalline size with the inclusion of nanoparticles (N_A and N_B) proved the existence of interactions among CH, TPP, and GA, resulting in a more amorphous structure [25]. According to Sujima-Anbu et al. [23], CH nanoparticles consist of a dense network structure of interpenetrated polymer chains crosslinked to each other by TPP.

Additionally, the analysis of peak diffraction located at $2\theta = 22^\circ$ evidenced that the lowest $N_{(A, B)}:CH$ ratio did not produce appreciable changes in the peak position. However, relations greater than $N_{(A, B)}:CH$ 1 : 1 showed a shift toward higher 2θ values (Figures 3(a) and 3(b)).

3.5. PALS Studies. For semicrystalline or amorphous materials, positrons will encounter various molecular environments such as free volume sites and crystalline defects [26]. The spectrum of a polymeric material can be fitted with three mean lifetimes, corresponding to the annihilation of p-Ps (τ_1, I_1), free positron (τ_2, I_2), and o-Ps (τ_3, I_3). Due to its sensitivity toward vacancies or defects, technique PALS has been applied to characterize CH nanocomposite materials.

As can be seen in Figure 4(a), the higher the N:CH ratio, the lower the T_g s obtained from the corresponding curves as was informed in *Thermal analysis section*. It is noteworthy that the molecular organization corresponded to an amorphous rubber state since the T_g of the systems remained above the room temperature. The inclusion of the nanoparticles caused an increase in lifetime, τ_3 , and, hence, in free volume that presented a maximum for N:CH ratio 1 : 1.

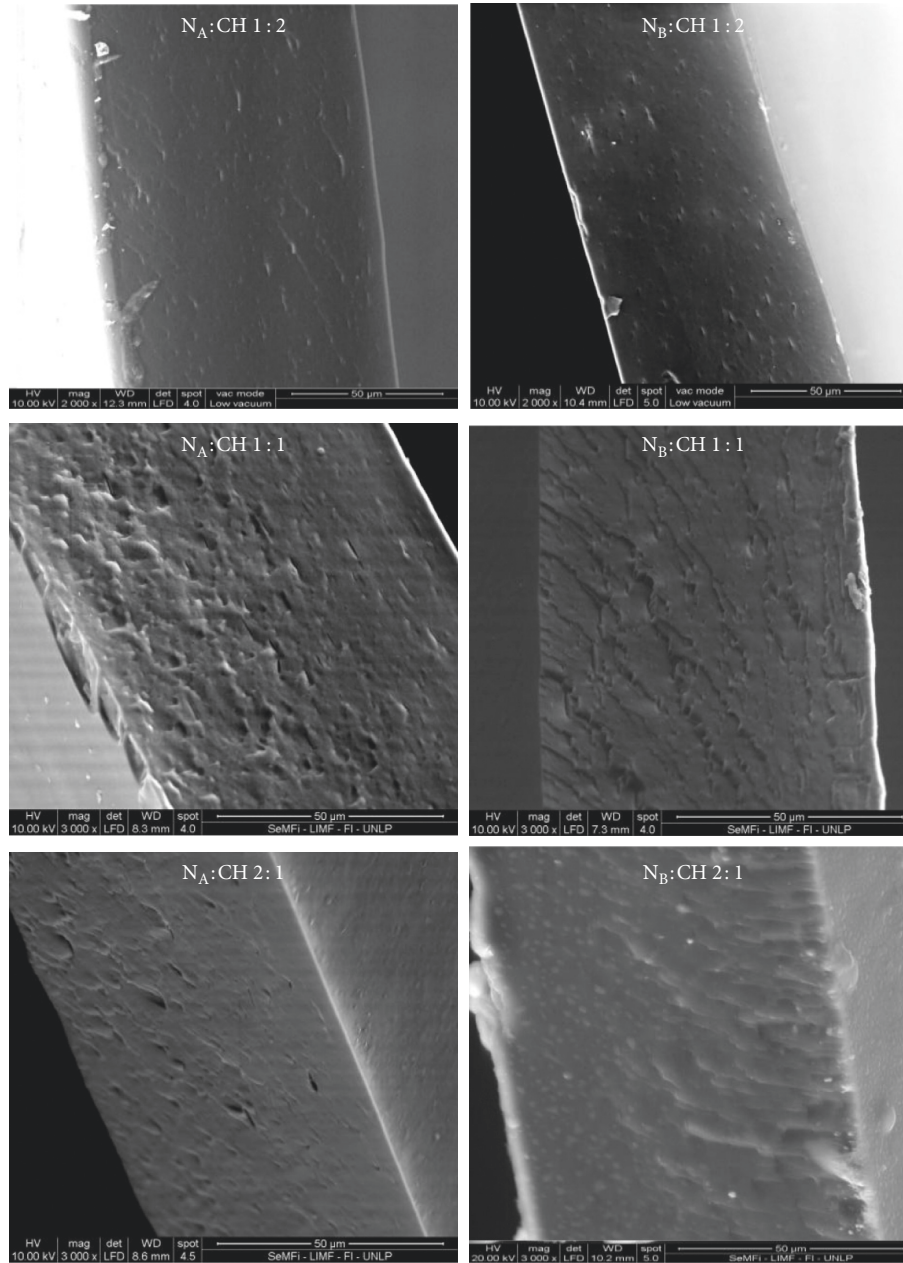


FIGURE 2: SEM cross-section micrographs of nanocomposite matrices with different $N_A:CH$ and $N_B:CH$ ratios. Scale bars are indicated in images.

For the same nanoparticle fraction, a parallelism between T_g values of the formulations containing either N_A or N_B as well as free volume values was observed. Additionally, the variables of $N_A:CH$ matrices remained above those obtained for $N_B:CH$ films.

It was not possible to infer the existence of a linear relation between both glass transition temperatures and the free volume irrespective of the $CH:TPP$ ratio (data not shown). This phenomenon may, in part, depend on the fact that the analysis must be based not only on o-Ps variables but also on free positron parameters.

Consequently, considering I_2 and τ_2 , which are associated with the annihilation of positrons in interfaces, a lineal

increasing relation between T_g s and I_2 and an inverse relationship between T_g s and τ_2 were observed (Figures 4(a) and 4(b)).

According to Sharma et al. [19], this lack of linearity of o-Ps lifetime could be explained considering the existence of poor dispersion (aggregation) with increasing nanoparticle concentration. The nanoparticles tend to form aggregates, which may be due to the large surface/volume ratio and intermolecular interactions, such as electrostatic, hydrophobic, and van der Waals interactions [27].

The increase in the intensity I_2 can be explained on the basis of the increase in concentration of the interfaces with the incorporation of a greater proportion of nanoparticles.

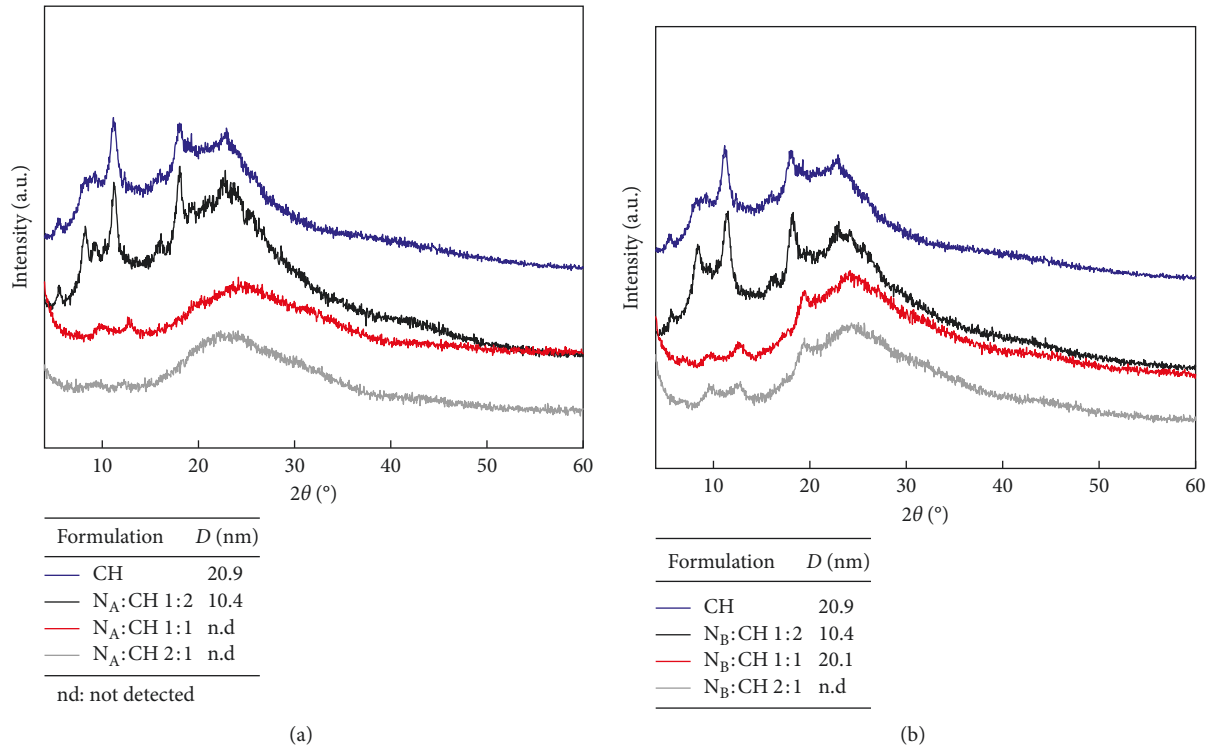


FIGURE 3: X-ray diffractograms of nanocomposites formulated with different N_A :CH (a) and N_B :CH ratios (b). Tables show the crystallite size estimated by using the Debye-Scherrer equation.

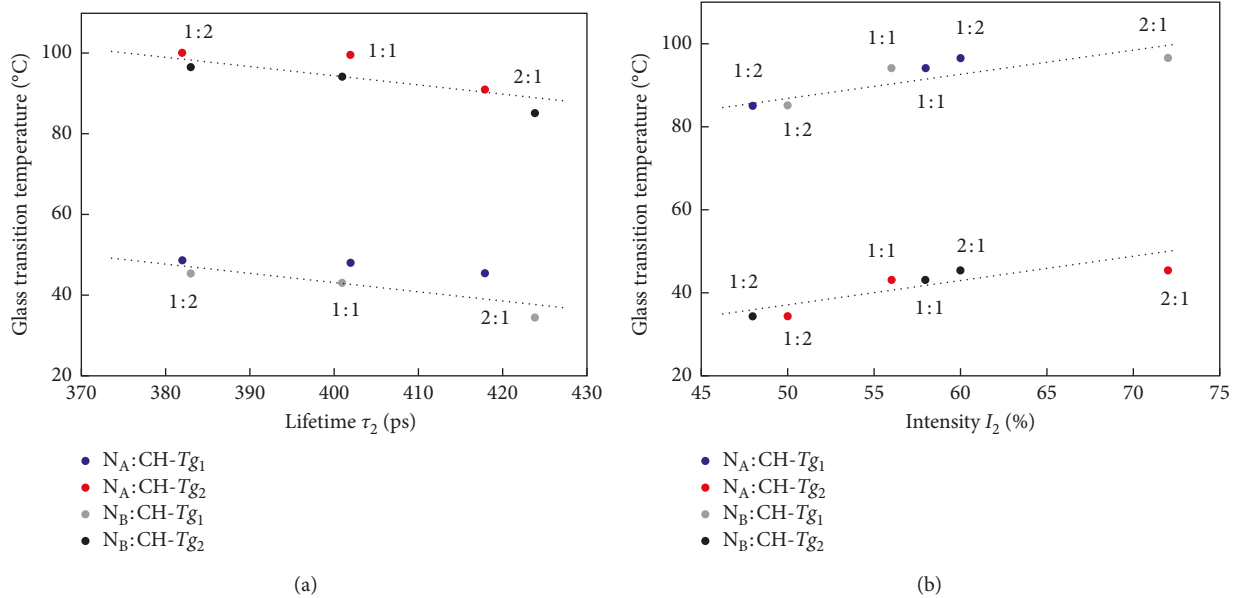


FIGURE 4: Correlation between (a) glass transitions temperatures (T_{g1} and T_{g2}) and o-Ps lifetime (τ_2) and (b) glass transitions temperatures (T_{g1} and T_{g2}) and the corresponding intensity I_2 (%) for different N_A :CH and N_B :CH ratios.

The positron trapping process in the defects present in the interfaces proved variations in both lifetime components, τ_2 and τ_3 [28].

The interfacial interaction among the CH polymer and the nanoparticles may restrict the chain motion, leading to a reduction in size of free volume in the vicinity of the

nanoparticles. As a result, o-Ps lifetime increased with the N:CH ratio increase.

3.6. ATR-FTIR Spectra. Nanocomposites spectra showed characteristic peaks at 1637 cm^{-1} (C-O stretching in amide

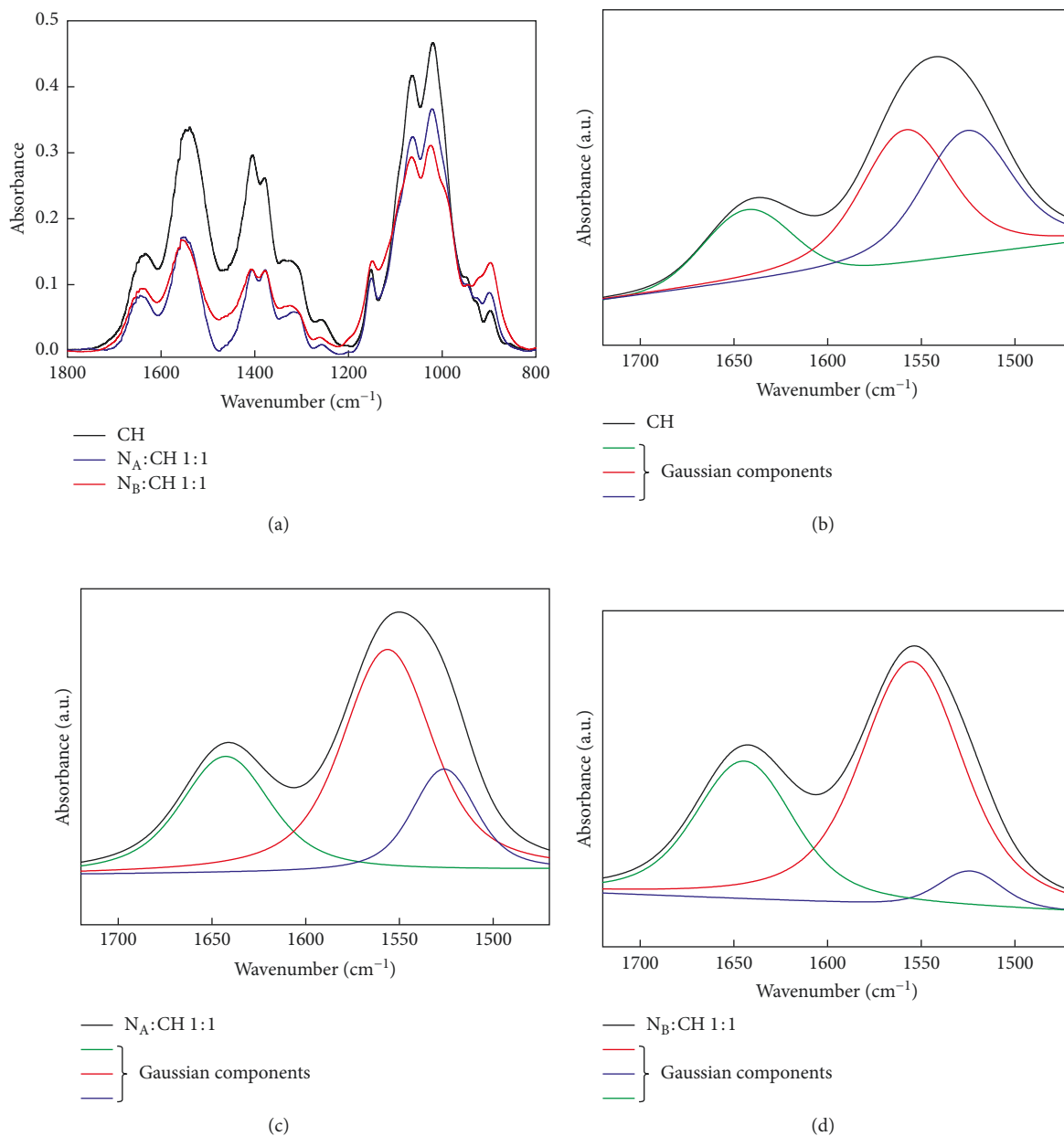


FIGURE 5: ATR-FTIR spectra of CH and nanocomposite matrices with 1:1 $N_{(A,B)}$:CH ratio showing the region 1800–800 cm^{-1} (a). Unresolved FTIR spectra and curve-fitting analysis in the amide I and amide II regions for CH (b), N_A :CH 1:1 (c), and N_B :CH 1:1 (d). Experimental data (black full lines) and individual Gaussian components (color full lines) are shown.

group, amide I vibration), 1538 cm^{-1} ($-\text{NH}_2$ bending in nonacetylated 2-aminoglucose primary amine), and 1377 cm^{-1} (C-O stretching of primary alcohol groups) (Figure 5). The absorption bands at 1150 cm^{-1} (antisymmetric stretching of the C-O-C-glycosidic linkage between chitosan monomers) and 1018 cm^{-1} (skeletal vibrations involving the C-O stretching) attributed to the chitosan saccharide structure were observed [29].

Figure 5(a) depicts the development of the band at 900 cm^{-1} as a result of the inclusion of nanoparticles (N_A and N_B) attributed to the antisymmetric stretching of the bending P-O-P of the TPP [30]. With the increase of the nanoparticle ratio, the peak intensity significantly increased,

and the shift of amide II band was visualized as a consequence of the interactions between the CH amino groups and the phosphate groups of the polyanion that produce an ionic crosslinking with the subsequent CH precipitation. These findings are analogous to those found by Hu et al. [31] and de Pinho Neves et al. [32].

The use of the spectra deconvolution method improved the resolution of the bands ascribed to amide I and II (Figures 5(b)–5(d)). The deconvolution of the band located at 1637 cm^{-1} showed a single peak after iterative curve-fitting procedure, whereas the amide II band was unfolded in two peaks located at 1555 and 1524 cm^{-1} . Ramon et al. [33] attributed these bands to intramolecular

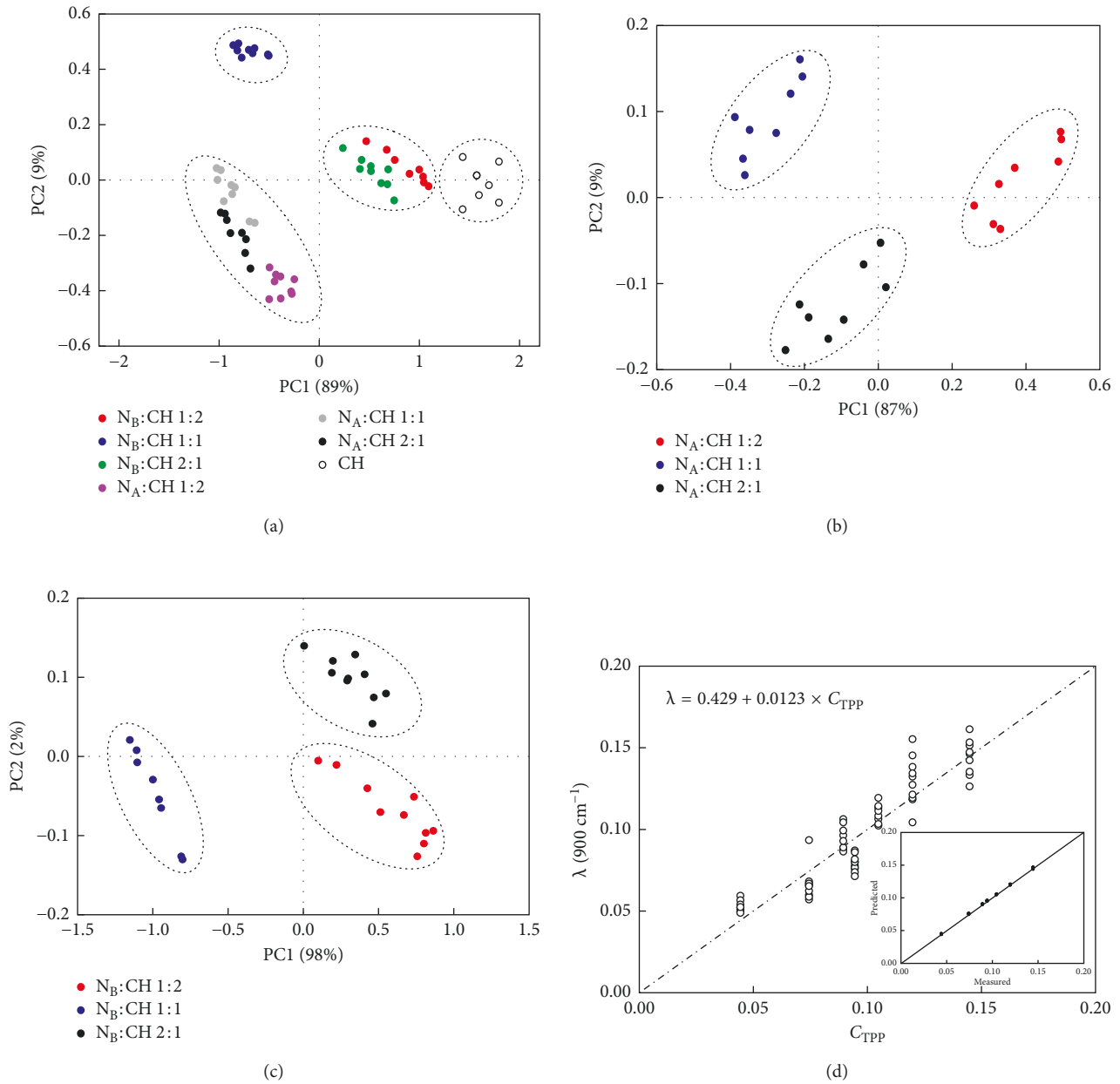


FIGURE 6: PCA biplot of CH and nanocomposite matrices in the ATR-FTIR region 1700–1400 cm^{-1} (a). PCA score plots show the cluster distribution of nanocomposites formulated with different N:CH ratios for N_A (b) and N_B (c). Correlation between the absorbance corresponding to 900 cm^{-1} and the concentration of tripolyphosphate (C_{TPP}) obtained from PLS model (d). Figure inserted in figure (d) shows predicted versus measured values.

hydrogen-bonded, intermolecular hydrogen-bonded, and free form of non-hydrogen-bonded subbands.

3.7. PCA and PLS Analysis. Principal components analysis (PCA) was carried out to better understand the interactions of the components of the nanocomposites. The PCA analysis involves a calculation of a covariance matrix of a data set to minimize redundancy and maximize variance. In a biplot, the length of the lines approximates the variances of the variables [34]. The systematic variation of the experimental values was mainly explained by means of 2 PCs. In

Figure 6(a), the first and second principal components (PC1 and PC2) represented 89% and 9% of the total variance, respectively.

Analyzing the biplot can be inferred that the incorporation of the nanoparticles produced a discrimination of the samples (CH, $N_A:CH$, and $N_B:CH$) into clusters when the ATR-FTIR region 1400–1750 cm^{-1} was considered. A separation between $N_A:CH$ and $N_B:CH$ formulations along PC1 axis was observed, but whereas all $N_A:CH$ nanocomposites were located at the left of PC2, $N_B:CH$ were positioned at both sides of PC2 axis. The high percentage of variance explained in PC1 axis (89%) indicated

the accuracy of ATR-FTIR to discriminate among the nanocomposite matrices. In particular, the high density of the cluster corresponding to $N_A:CH$ indicated the low variability among these samples (Figure 6(a)). The high overlapping of spectra corresponding to all $N_A:CH$ formulations is very graphic in this respect. On the contrary, the highest variability of spectra corresponded to $N_B:CH$ 1:1 (Figure 6(a)).

Although this preliminary analysis was not able to underlined among all formulations, a posteriori analysis carried out on either $N_A:CH$ or $N_B:CH$ separately, allowed a discrimination within each of these nanocomposites in new clusters, as can be seen in Figures 6(b) and 6(c). Similar results (not shown) were obtained when the analysis was conducted on the FTIR region $2600\text{--}3600\text{ cm}^{-1}$.

Additionally, the analysis carried out by the PLS method allowed the adjustment of the experimental data of a selected wavenumber (900 cm^{-1}) as a function of the TPP concentration (C_{TPP}) by using the model inset in Figure 6(d). This peak located at 900 cm^{-1} was selected as a band representative of the modifications experimented by the CH matrix as a consequence of the incorporation of the nanoparticles. A satisfactory fit between measured and predicted values was observed (inset in Figure 6(d)).

4. Conclusions

The nanocomposites formulated with the highest nanoparticle proportion and the highest TPP concentration achieved significant modifications on the CH matrices, by virtue of the crosslinking of the biopolymer with GA and TPP, specially for N:CH ratio 1 : 1. A further increase in the N:CH ratio higher than 1 : 1 produced no significant improvement in the nanocomposite properties.

The results obtained by means of MDSC, ATR-FTIR, PALS, and SEM techniques showed an interaction between CH and GA, owing to the encapsulation of the active compound. The PALS technique was adequate to detect the modifications of both size and concentration of the free volume as a consequence of different N:CH and CH:TPP ratios. In addition, this procedure showed a good correlation between glass transition temperatures (T_g) obtained by using the MDSC technique and the PALS parameters τ_2 and I_2 ascribed to the annihilation of positrons in the interface.

In conclusion, the structural insight revealed the changes in nanocomposites modifying the performance of the materials. In this sense, further studies are needed concerning future applications in food and agricultural fields.

Conflicts of Interest

The authors declare that they have no conflicts of interest.

Acknowledgments

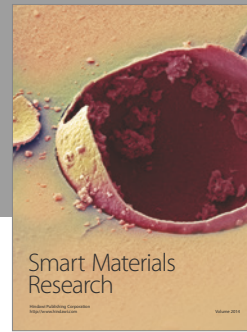
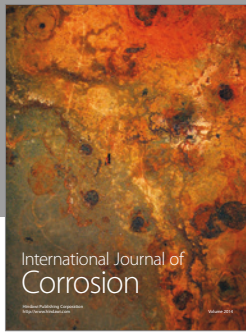
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