CHEMISTRY OF MATERIALS

Photoluminiscent 1−2 nm Sized Silicon Nanoparticles: A Surface-Dependent System

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Ⓢ *Supporting Information*

ABSTRACT: The effect of derivatization and temperature on the photoluminescence of 1−2 nm size silicon particles of different origin is investigated in an attempt to understand the effect of surface on the particles' photoluminescence. To this purpose, silicon nanoparticles were synthesized by electro-chemical (top-down) and wet chemical (bottom-up) procedures. Further derivatization by silylation or sylanization yielded particles with $\equiv$Si—C$\equiv$, $\equiv$Si—O—Si$\equiv$, and $\equiv$Si—O—C$\equiv$ groups at the interface. A detailed analysis of the corresponding excitation−emission matrices strongly indicates that different surface atomic arrangements contribute to the energy gap. In particular, particles with $\equiv$Si—O—Si$\equiv$ groups at the interface show photoluminescence independent of the crystalline structure and on their further surface derivatization with different organic molecules. The lifetime and spectrum shape of all synthesized particles are invariable to changes in temperature in the range 270−330 K despite a significant reduction in the photoluminescence intensity being observed with increasing temperature; such behavior supports a thermal equilibrium between dark and bright conformations of the particles. The observed results are of importance for optimizing the use of silicon nanoparticles as optical sensors and therapeutic agents in biological systems.

KEYWORDS: surface states, thermal quenching, Si/SiO$_2$ interface, optical properties, surface chemistry
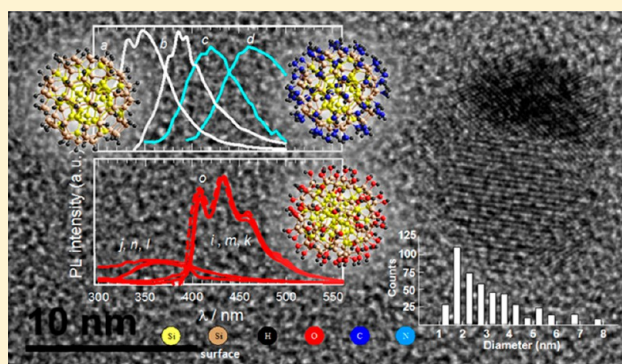
■ INTRODUCTION

Silicon nanoparticles of 1−5 nm size (SiNPs) have received great attention as they combine photoluminescence emission (PL) spanning the visible and near-IR spectral region, with the capability for singlet oxygen ($^1O_2$) and superoxide ($O_2^{\bullet-}$) generation.[1,2] Remarkable properties of nanostructured silicon are their biocompatibility, biodegradability, and tunable surface derivatization for drug delivery.[3,4] SiNPs surface derivatization with active functional groups, biomolecules, and biocompatible polymers is a key step in the development of their many technological applications. An adequate selection of the surface groups renders the particles' special properties as dispersibility in the desired medium, surface protection against oxidation, specific targeting, and opsonin association prevention.[5]

Work with SiNPs has shown that crystallinity, surface structure, and surface chemistry play a role in their photo-physical behavior. Interpreting and classifying the large number of energy gaps (EG) and Stokes shifts (SS) reported for SiNPs is difficult due to long-standing divergences in the synthesis of monodisperse samples with well-characterized surfaces, as a variety of groups (e.g., $\equiv$Si—OH, >Si$=$O, $\equiv$Si—O—Si$\equiv$,

$\equiv$Si—H, $\equiv$Si—C$\equiv$, etc.) present in different proportions may form the interface of the designed product. In fact, the fundamental photophysics of SiNPs is not currently as well understood as that of direct gap semiconductors such as CdSe.[6]

Electron density functional theory assuming finite barrier for the potential confinement applied to H-saturated silicon clusters predicts smooth variations of the EG with size.[7] Increased EG observed experimentally for crystalline clusters were suggested to be a consequence of symmetry and reduced angular distortion in the Si—Si bonds.[8] These authors predicted that localization of the HOMO and the LUMO occurred throughout the core for unreconstructed surfaces (e.g., Si$_{35}$H$_{36}$ bulk surface) whereas localization at the surface was observed for reconstructed surfaces (e.g., Si$_{29}$H$_{24}$). Moreover, experimental and computational work with SiNPs demonstrated the fundamental importance of surface chemistry in dictating their photophysical behavior.[6,9,10] In particular,

single bonded Si—OH and Si—O—Si groups on Si$_{29}$ and Si$_{47}$ cores have relatively slight influences on the band edges and gap, whereas double bonded Si=O surface structures significantly affect them. The increase in surface coverage causes a further change of the EG.[9] More recent studies suggest that Si—O—Si structures on the curved surface of <3 nm size SiNPs, but not on the facet of >3—4 nm particles, lead to localized levels in the EG. As a consequence, a red-shift in the PL spectra of small oxidized SiNPs is observed.[11]

The temperature dependence of the PL of SiNPs is suggested to be an intrinsic property of the particles strongly depending on size.[12−14] SiNPs from different synthesis routes are reported to show increased PL intensity with decreasing temperature until ca. 60−80 K. A further decrease in temperature rapidly diminishes the PL intensity along with an increased PL lifetime. These observations are discussed in terms of the competition between temperature-dependent radiative (k_r) and nonradiative ($\sum_i k_{ni}$) decay rates of the highly localized excitons, as those provided by defects in the nanocrystal or at the Si/SiO$_2$ interface.

Herein we investigate the effect of surface derivatization and temperature on the PL of SiNPs of different origin in an attempt to understand the effect of surface on the particles PL. This knowledge is necessary for optimizing the use of SiNPs as optical sensors and therapeutic agents in biological systems. To that purpose, SiNPs were synthesized by electrochemical (top-down, TD) and wet chemical procedures (bottom-up, BU), and further derivatized by silylation or sylanization to yield particles with —NH$_2$ terminal groups and different chemistry at the interface (≡Si—C≡, ≡Si—O—Si≡, and/or ≡Si—O—C≡).

■ EXPERIMENTAL SECTION

Materials. Toluene (99.7%, H$_2$O 0.005%), methyl 2-methylprop-2-enoate, methanol, cyclohexane, 9,10- diphenylanthracene, LiAlH$_4$ (95%), SiCl$_4$ (99%), ethyl ether (p.a. 99.9%), tetraoctylammonium bromide (98%), (3-aminopropyl)triethoxysilane, and 2-propen 1-amine were purchased from Sigma-Aldrich and employed without further purification. Deionized water (>18 MΩ cm, <20 ppb of organic carbon) was obtained with a Millipore system. Nitrogen (4 bands quality) and oxygen gas were both from La Oxígena S.A., Argentina.

Equipment. FTIR spectra were obtained with a Bruker EQUINOX spectrometer. Either KBr disks or silicon wafers were used as holders. Spectra were taken in the 4000−400 cm^{-1} range with 1 cm^{-1} resolution. Attenuated total reflection infrared (ATR-IR) spectroscopy used for the analysis of the surface of derivatized particles was performed with a Nicolet 380 FTIR ATR accessory using a ZnSe crystal and an incident beam angle of 45°.

The attenuance spectra were recorded with a double-beam Shimadzu UV-1800 spectrophotometer in a 1 cm quartz cuvette at a scan rate of 300 nm/min. The nanoparticle light scattering is calculated by fitting the 450−800 nm range attenuance to a $a \times \lambda^{-4}$ dependence. The corrected absorbance spectrum is obtained by subtraction of the scattering from the measured attenuance spectrum.

PL measurements were performed with a Jobin-Yvon Spex Fluorolog FL3-11 spectrometer. The fluorimeter is equipped with a Xe lamp as the excitation source, a monochromator with 1 nm bandpass gap for selecting the excitation and emission wavelengths, and a red sensitive R928 p.m. as detector. All spectra were corrected for the wavelength-dependent sensitivity of the detector and the source. Additionally, emission spectra were corrected for Raman scattering by using the solvent emission spectrum. To estimate the emission quantum yield (Φ), emission spectra were collected at various excitation wavelengths. Identical measurements (excitation conditions, lamp energy, and spectrometer band-pass) were performed on 9,10-diphenylanthracene in cyclohexane.[2,15,16] The temperature

was controlled to ±0.1 °C with an F-3004 Peltier sample cooler controlled by a LFI-3751 temperature controller (Wavelength electronics).

Time-resolved PL measurements were performed with either a Jobin-Yvon Spex Fluorolog FL3-11 spectrometer (vide supra) equipped with TCSPC with LED excitation at 341 and 388 nm (fwhm ∼ 400 ps) or a mode-locked Ti:Sapphire laser (Spectra Physics, Millennia-pumped Tsunami) with a 130-fs pulse duration operating at 80 MHz. The laser output was sent through a frequency doubler and pulse selector (Spectra Physics Model 3980) to obtain 370−450 nm pulses at 4 MHz. Fluorescence emission was detected at the magic angle using a double grating monochromator (Jobin Yvon Gemini-180) and a microchannel plate PMT (Hamamatsu R3809U-50). The instrument response function was 35−55 ps. Data was acquired by a Becker-Hickl SPC-830 card. Temperature was controlled to ±0.1 °C with an Oxford Optistat DN liquid nitrogen-cooled optical cryostat equipped with an Oxford ITC 601 controller. Data was globally fitted as sum of exponentials including IRF deconvolution until optimal values of χ^2, residuals, and standard deviation parameters were attained.

TEM micrographs were taken with a JEOL 2010 F microscope using a carbon-coated 300-mesh copper grid. Images were analyzed employing the Image Tool 3.0 software (Health Science Center of the University of Texas, San Antonio, TX, USA). Particle diameters were determined assuming that the particle area obtained from the TEM images is the projection of a spherical particle. A log-normal function was found to best describe the size distribution of the 500 particles sample.

The X-ray photoelectron spectroscopy (XPS) spectra were obtained under UHV with a XR50 Specs GmbH spectrometer with Mg Kα as the excitation source and a PHOIBOS 100 half sphere energy analyzer. Internal calibration was performed with Au 4f$_{7/2}$ (binding energy, BE, 84.00 eV) and with the C 1s peak at BE = 284.6 eV due to adventitious carbon. Also, to get a better insight into the chemical environment of the different atoms, high resolution XPS spectra were taken and resolved by Gaussian−Lorentzian fitting, keeping χ^2 values between 1 and 1.25.

Bilinear Regression Analysis. For low fluorophore absorbance (<0.05), the steady state emission intensity at the given emission wavelength, $I(\lambda_{em})$, is proportional to the absorption coefficient of the fluorophore at the excitation wavelength, $\varepsilon(\lambda_{ex})$, the concentration of the fluorophore, and a factor $F(\lambda_{em})$ reflecting the shape of the emission spectrum. The bilinear regression analysis takes advantage of the linearity of $I(\lambda_{em})$ with both $\varepsilon(\lambda_{ex})$ and $F(\lambda_{em})$ to retrieve information from the experimental excitation−emission matrix on the minimum number of species and on their relative emission and absorption spectra.[16] It should be noted that the recovery of the excitation−emission spectrum of all the individual components from a bilinear analysis of a multicomponent system is not always possible in those cases where the emission and excitation spectra of the different species are strongly superimposed.[17] A further disadvantage is a low contribution of one of the species to the overall emission. In these cases, the different components may not be differentiated by a bilinear analysis, and the individual recovered species are rather a family of strongly optically related compounds.

SiNPs Synthesis. TD approaches involved an adaptation of the electrochemical method with HF etching of porous Si.[1] Briefly, clean crystalline Si wafers (p-type, with resistivity between 1.0 and 10.0 Ω cm) were electrochemically etched (20 mA electrical current) in a Teflon cell containing the electrolytic solution (16% HF, 16% CH$_3$OH v/v). After anodization, the wafer was rinsed with methanol and toluene, further immersed in Ar-saturated toluene, and Si-NP released by immersion in an ultrasound bath for 2 h (Ney Dental Inc.; 50/60 Hz; 100 W). The resulting suspension was purified by filtration through hydrophilic 100 nm PVDF membrane filters (MILLEX-HV, Millipore), and both the solid residue and the filtrate were separated to obtain different SiNPs.

The solid residue was suspended in a 3% HF methanol solution, sonicated, and stored in the dark for two hours to break up and etch agglomerates.[18] After water addition, the well-dispersed aqueous

Scheme 1. Summary Flow Chart on the Synthesis Pathways Followed for Obtaining the Different SiNPs[a]
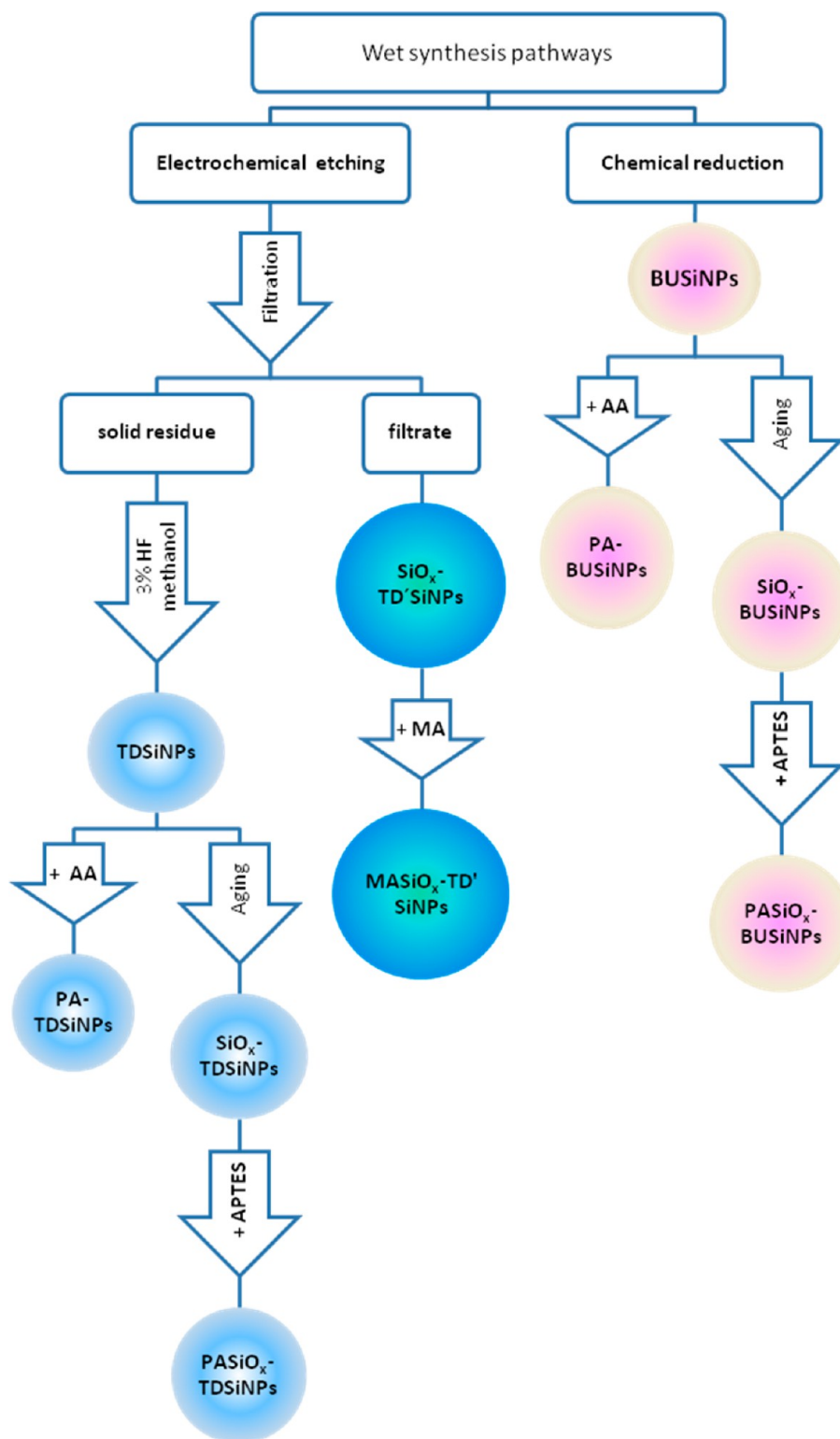
[a]AA, MA, and APTES stand for 2-propen 1-amine, methyl 2-methylprop-2-enoate, and (3- aminopropyl)triethoxysilane, respectively.

suspension of SiNPs was shaken with cyclohexane to extract the hydrophobic nanoparticles with H-passivated surfaces to the organic phase. The solvent was evaporated and the remaining silicon particles dispersed in toluene for further use. The "as obtained" particles will be referred to as TDSiNPs. When required, TDSiNPs were allowed to age upon standing in an air-saturated suspension in the dark for several weeks, leading to the formation of surface-oxidized particles, SiO$_x$-TDSiNPs. Surface-oxidized 3 nm size SiNPs with a SiO$_x$ layer and Si—H surface functionalities are obtained from the filtrated suspension.[1] Because of their common TD synthetic procedure, these particles are referred to as SiO$_x$-TD′SiNPs.

The BU-approach synthesis involved an adaptation of the LiAlH$_4$ reduction of SiCl$_4$ in the presence of tetraoctylammonium bromide reversed micelles.[16] Freshly prepared particles showing H- and Cl-

Scheme 2. Surface Composition of TDSiNPs Derivatized with 2-Propen 1-Amine and (3- Aminopropyl)triethoxysilane and TD′SiNPs derivatized with Methyl 2-Methylprop-2-enoate[a]
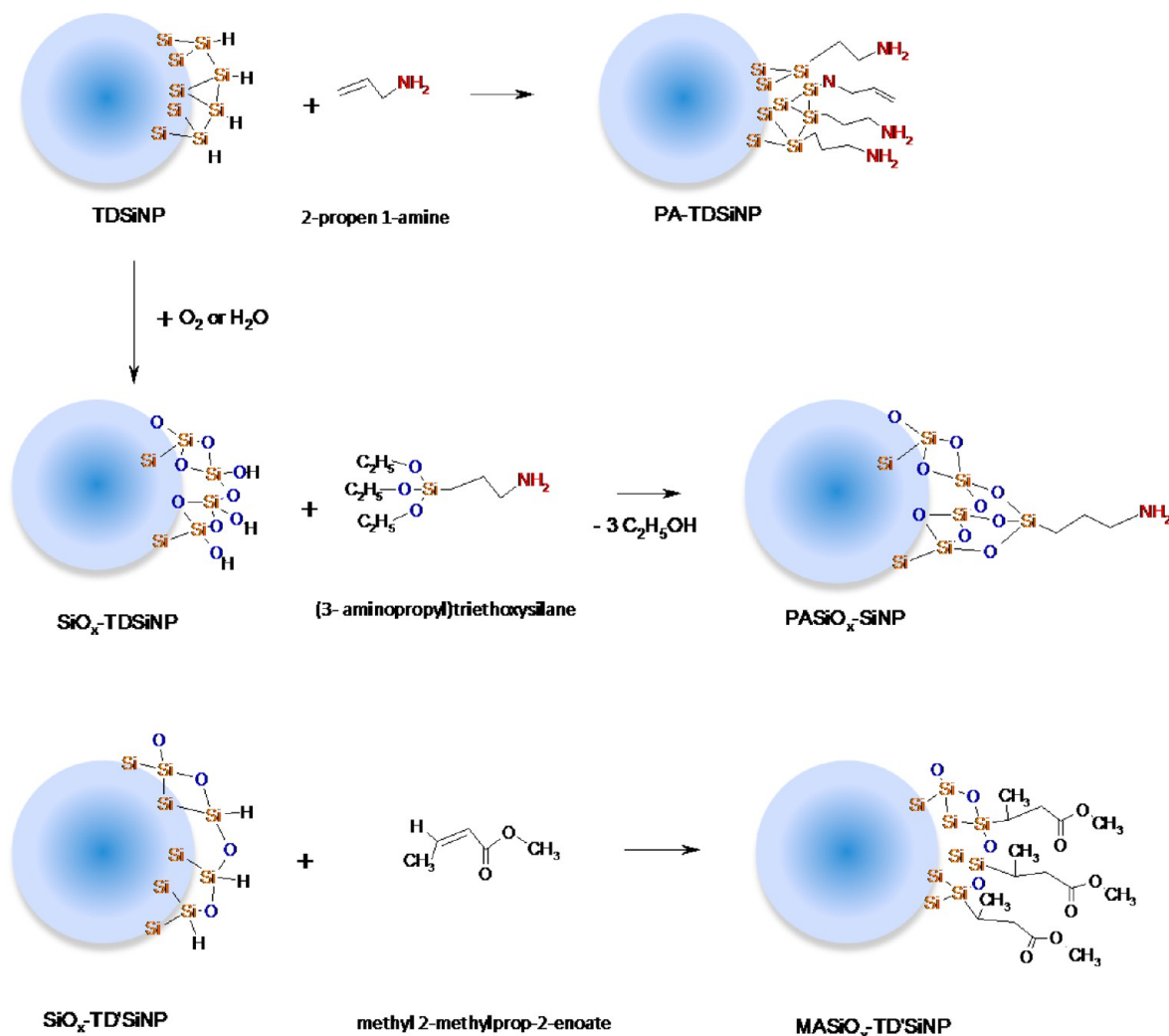
[a]The blue circles stand for the silicon core of the particles.

passivated surfaces are named BUSiNPs. These surface groups are oxidized to Si—OH during storage under air-saturation for two weeks, as confirmed by the appearance of a broad band near 3400 cm^{-1} in the FTIR spectrum, characteristic of Si—O—H stretching vibration. Aged particles are denoted as SiO$_x$-BUSiNPs.

Surface Derivatization. SiNPs were capped with organic molecules either by photoinitiated silylation or by thermal silanization. The silylation process involved mixing of a colloid dispersion of either fresh TDSiNPs or BUSiNPs in toluene with microliter quantities of 2-propen 1-amine (reactant in defect). The resulting suspension was irradiated for 5 h with 254 nm light from eight Rayonet Lamps (RPR2537A, Southerm N.E. UltravioletCo.) to yield PA-TDSiNPs and PA-BUSiNPs. A similar procedure using 350 nm light from a Rayonet Lamp (RPR3500A, Southerm N.E. UltravioletCo.) was applied for freshly prepared methyl 2-methylprop-2-enoate derivatized SiO$_x$-TD′SiNPs containing surface Si—H bonds (MASiO$_x$-TD′SiNPs).

The silanization procedure involved coupling of (3-aminopropyl)-triethoxysilane to terminal Si—OH groups.[18,19] Therefore, suspended surface-oxidized BUSiNP and TDSiNP in toluene were treated with (3-aminopropyl)triethoxysilane and stirred at room temperature for 24 h. Removal of excess organosilane and tetraoctylammonium bromide was achieved by membrane dialysis. Oxidized particles derivatized with (3- aminopropyl)triethoxysilane are referred to as PASiO$_x$-BUSiNPs and PASiO$_x$-TDSiNPs.

For clarity, a flowchart summary depicting the synthesis pathways is given in Scheme 1.

■ RESULTS AND DISCUSSION

Silicon nanoparticles were obtained in the laboratory by top-down (TD) and bottom-up (BU) approaches, yielding H-passivated crystalline TDSiNPs and amorphous BUSiNPs with Si—H and Si—Cl surface groups, respectively (vide supra). These particles were either derivatized to cover the surface with propylamine terminal groups to yield PA-TDSiNPs and PA-BUSiNPs or allowed to age to yield surface oxidized SiO$_x$-TDSiNPs and SiO$_x$-BUSiNPs. The latter particles were further derivatized to yield propylamine-terminated PASiO$_x$-TDSiNPs and PASiO$_x$-BUSiNPs, respectively. SiO$_x$-TD′SiNPs particles were obtained by a modification of the TD approach involving the simultaneous oxidation and etching of porous silicon.[1] They differ from SiO$_x$-TDSiNPs in that they present Si—H, Si—O—Si, and Si—OH structures in their surface. SiO$_x$-TD′SiNPs particles were further derivatized to yield methyl 1-methyl-propanoate-terminated MASiO$_x$-TD′SiNPs. Scheme 2 summarizes the derivatization reactions for TDSiNPs and shows the average surface composition of the particles.
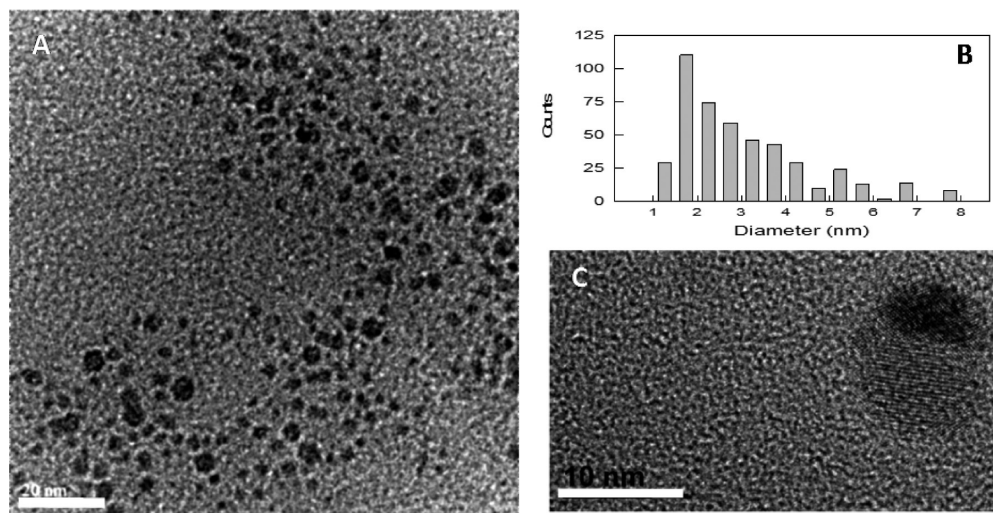
Figure 1. (A) TEM image of TDSiNPs, (B) associated size histogram of 500 non-touching particles, and (C) HRTEM micrograph of agglomerated particles showing silicon d_{111} crystalline planes.
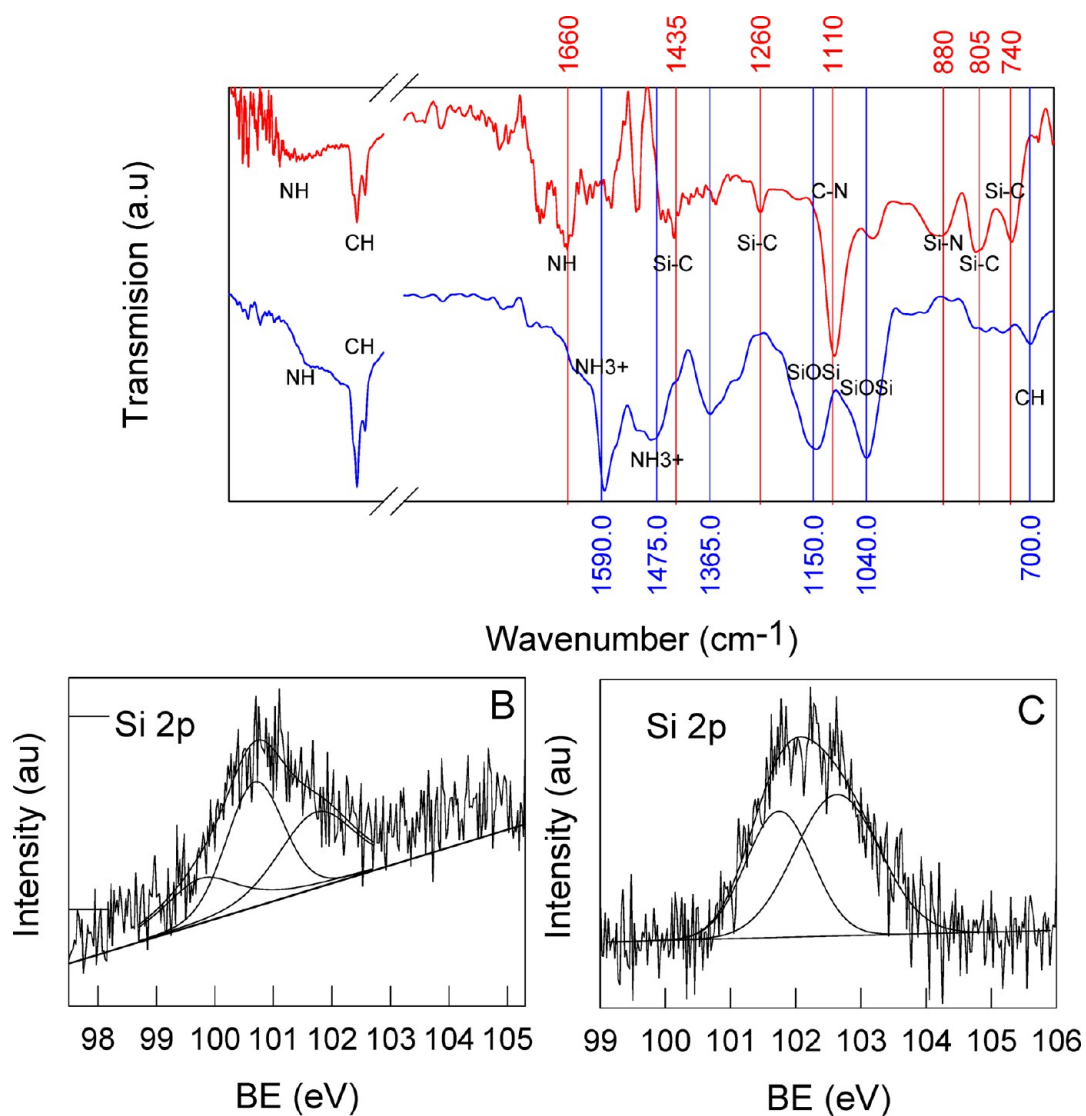

Figure 2. (A) FTIR spectra of PA-TDSiNPs (red spectrum) and PASiO$_x$-TDSiNPs (blue spectrum), deposited from toluene suspensions on Si wafer supports. The vertical solid lines stand as guides to the eye for the peak positions. (B) and (C) stand for the high resolution Si 2p XPS peak of PA-TDSiNPs and PA-SiO$_x$TDSiNPs, respectively.

Particle Characterization. An average size of 1.8 nm with a high dispersion toward bigger sizes is observed for TDSiNPs from the size histogram, as shown in Figure 1A,B. However, a thorough analysis of the micrographs (Figure 1C) indicates important particle agglomeration. In fact, atomic force microscopy performed on mercaptopropyl-derivatized TDSiNPs after drop-casting on a mica support and the heights extracted from scanning tunneling microscopy topography of mercaptopropyl-terminated TDSiNPs on a Au(111) substrate show size histograms indicating highly monodispersed samples with a mean particle height of (1.7 ± 0.3) and around (1.6 ± 0.5) nm, respectively.[20] The image in Figure 1C and other HRTEM micrographs shown in S.I. Characterization (Supporting Information) clearly show Si crystalline planes with lattice spacing of 3.1 ± 0.1 Å, in agreement with that reported for the d_{111} plane of SiNPs.[21,22] On the other hand, aged BUSiNPs showed an average particle size of (1.7 ± 0.8) nm and no evidence of a crystalline structure could be found[16] (see also TEM micrographs in S.I. Characterization, Supporting Information).

The analysis of the XPS and IR spectra of the different particles allows the determination of the average surface composition. Figure 2A, top spectrum, shows the FTIR spectrum of PA-TDSiNPs. The observed N—H vibrations in the 3400 (weak), 1650−1580, and 910−665 cm^{-1} regions are characteristic of primary and secondary amines. The C—N stretching vibration of aliphatic amines is observed at 1260 cm^{-1}. The presence of bands at 1435, 1270, 805, and 740 cm^{-1} due to Si—C symmetric and asymmetric deformation, stretching, and CH_2 rocking in Si—CH_2 respectively, as well as peaks in the 2970−2880 cm^{-1} region due to CH_2 stretching and bending, confirm PA bonding to TDSiNPs.[23,24] The corresponding XPS spectrum shows Si 2p signals, Figure 2B, at 99.70 (10%), 100.7 (57%), and 101.8 eV (33%), characteristic of Si^0, Si—C, and Si—N environments, respectively.[25−27] The N 1s signal may be described by a main peak at 399.4 eV characteristic of C—NH_2 groups. No significant contribution due to N in Si environments (peak at 398 eV[27]) is observed. However, the FTIR absorption at 880 cm^{-1} could be assigned to the presence of Si—N vibrations.[28] N-bonding to Si may be a secondary product of the radical mechanism involved in the synthesis reaction, as described in S.I. Scheme 1 (Supporting Information).

The FTIR spectra of PASiO$_x$-TDSiNPs, see bottom spectrum in Figure 2A, shows intense bands at 1150 and 1040 cm^{-1} characteristic of Si—O—Si in different environments, thus denoting surface-oxidized particles. Absorption bands in the region 1550−1450 cm^{-1} may be assigned to N—H vibrations in protonated amines.[29] The peak at 1332 cm^{-1} is assigned to CO_2 trapped by surface NH_3^+ in coincidence with that reported for protonated (3-aminopropyl)triethoxysilane films.[30] On the other hand, the Si 2p XPS peak displays the contribution of silicon environments with binding energies of 101.7 and 102.6 eV which may be assigned to Si(O—)$_2$ and Si(O—)$_3$—C environments,[31,32] in agreement with the expected PASiO$_x$-TDSiNPs surface depicted in Scheme 2. N 1s peaks at 399.4 eV may be assigned to C—NH_2 environments, and the IR peaks due to CH_2 stretching and bending vibrations in the 2970−2880 cm^{-1} and 1465−1440 cm^{-1} region further support the covalent link of organic moieties to the silicon framework. The absence of Si 2p XPS peaks around 104 eV and IR bands around 1250 cm^{-1} due to Si=O in bulk SiO_2 supports the negligible formation of SiO_2 structures.[33]

Ratios between Si, O, C, and N signals in the XPS spectrum corrected for the instrument sensitivity yield the average surface composition for the different SiNPs, as shown in Table 1 and in

Table 1. Average Size and Surface Composition of Solid Samples Obtained from the Different Synthesis and Derivatization Procedures[a]

particle	size, nm	average surface composition[b]
TDSiNPs	1.7 ± 0.3	-----
PA-TDSiNPs	1.7 ± 0.3	$Si_2C_5N_{4.4}H_y$
SiO$_x$-TDSiNPs	1.7 ± 0.3	-------
PASiO$_x$-TDSiNPs	1.7 ± 0.3	$Si_{12}C_{3.5}H_yN_1O_x$
MASiO$_x$-TD′SiNPs	3 ± 1[c]	$Si_3O_6(C_5O_2H_y)_1$[c]
BUSiNPs	1.7 ± 0.8	-----
PA-BUSiNPs	1.7 ± 0.8	-------
SiO$_x$-BUSiNPs	1.7 ± 0.8	$Si_{4.2}O_{3.8}ClH_y$
PASiO$_x$-BUSiNPs	1.7 ± 0.8	------

[a]Surface oxidized particles are in bold. [b]H "y" proportion was not determined though observed. [c]Data from ref 1.

agreement with the expected surface structures depicted in Scheme 2. FTIR and XPS characterization of BUSiNPs, PA-BUSiNPs, SiO$_x$-BUSiNPs, and PASiO$_x$-BUSiNPs also supports the expected surface structure, in agreement with the FTIR and XPS peak assignments observed for TDSiNPs derivatives. For further details refer to S.I. Characterization (Supporting Information). It should be recalled that PA-TDSiNPs also shows a small FTIR absorption at 850 cm^{-1} and peak contributions to the Si 2p band at 101.5 and 102.2 eV, characteristic of Si(C$_x$N$_y$) and Si—N environments.[25,28] In conclusion, PA derivatization of H-passivated SiNPs by a photoinitiated radical mechanism leads invariably to N-bonding to Si independently of the crystalline structure of the silicon network.

Photoluminescence Experiments. The PL of the different particles in Ar-saturated toluene suspensions shows particular features strongly depending on surface groups and synthesis procedures. Moreover, the dependence of the emission spectrum with the excitation wavelength for a given suspension indicates the contribution of particles of different size, morphology, surface chemistry, and/or oxidation grade within the sample.[16] The contribution of different emitters to the overall excitation−emission matrix was analyzed on the assumption that single chromophores existing in a unique form show excitation wavelength-invariant emission spectrum.[16] Figure 3 shows the emission and excitation spectra of contributing emitters as obtained from the bilinear analysis of the excitation−emission matrix of argon-saturated toluene suspensions of the different synthesized particles.

Decay times (τ) and quantum yields at 298 K (Φ^{298K}) were measured at selected excitation and emission wavelengths to focus on a particular contributing emitting species. Measured τ values for TDSiNPs and BUSiNPs and their surface-functionalized derivatives are within 1−7 ns, as depicted in Table 2. Under conditions of low excitation power density, the excited states of the particles decay unimolecularly and the PL quantum yield is $\Phi = k_r/(k_r + \sum_i k_{ni})$, where k_r stands for the radiative decay rate and $\sum_i k_{ni}$ for nonradiative processes. Therefore, considering the luminescence lifetime $\tau = (k_r + \sum_i k_{ni})^{-1}$, the product $\Phi \times \tau^{-1}$ yields k_r. Considering the data in Table 2,

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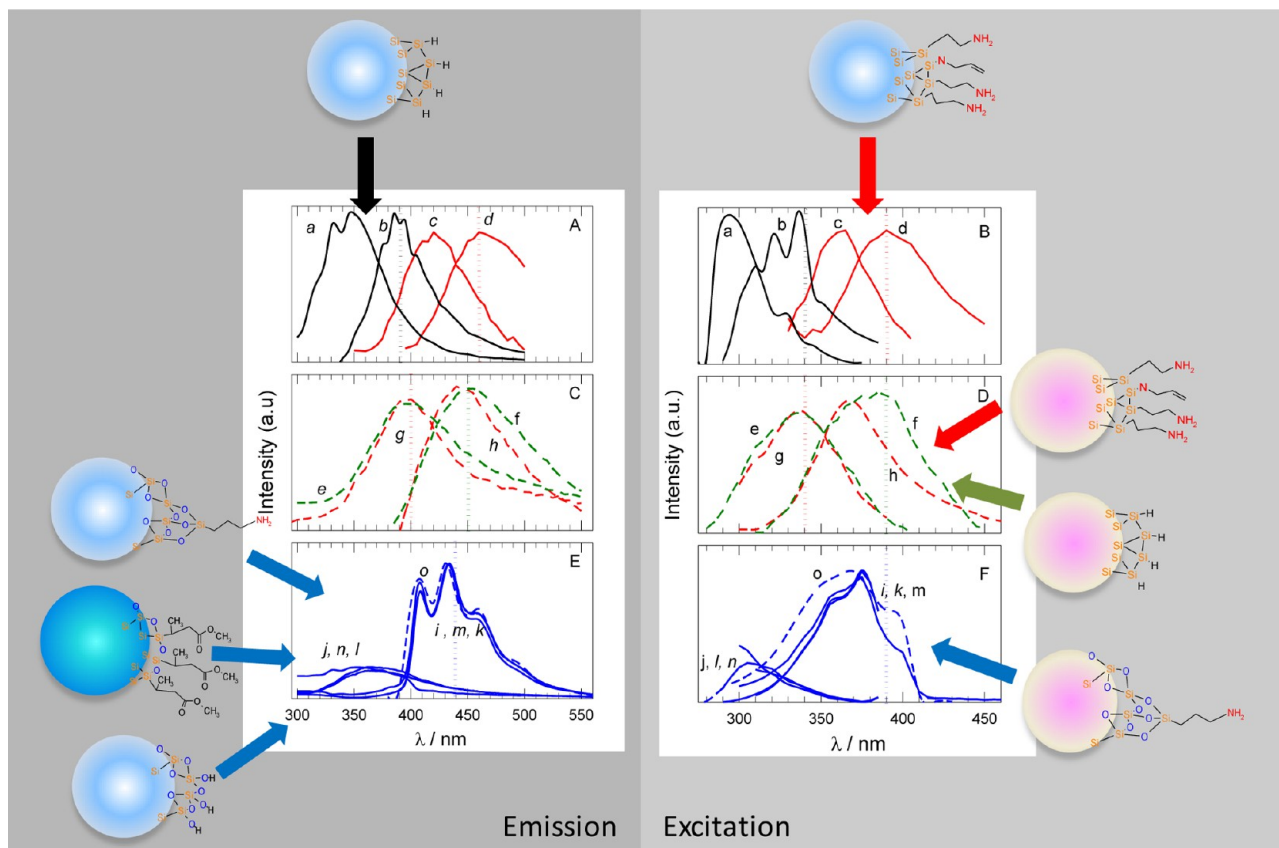dx.doi.org/10.1021/cm401666a | Chem. Mater. XXXX, XXX, XXX−XXX

Figure 3. Emission (A, C, and E) and excitation (B, D, and F) spectra at 313 K of argon-saturated toluene suspensions of contributing emitters as obtained from the bilinear analysis of the excitation–emission matrix of TDSiNPs (*a* and *b*, black solid lines), PA-TDSiNPs (*c* and *d*, red solid lines), BUSiNPs (*e* and *f*, green dashed lines), PA-BUSiNPs (*g* and *h*, red dashed lines), SiO$_x$-TDSiNPs (*i* and *j*, blue solid lines), PASiO$_x$-TDSiNPs (*k* and *l*, blue solid lines), MASiO$_x$-TD′SiNPs (*m* and *n*, blue solid lines), and PASiO$_x$-BUSiNPs (*o*, blue dashed lines). The vertical lines stand for the excitation (B, D, and F) and detection (A, C, and E) wavelengths in time-resolved experiments.

Table 2. Energy gap (EG), maximum emission (λ_{em}^{max}), Stokes shifts (SS), decay times (τ), and quantum yields (Φ) of the different emitting particles obtained at 298 K in argon-saturated suspensions[a]

particle	emitter	EG (± 0.05), eV	λ_{em}^{max}, eV	SS (± 0.05), eV	τ(%), ns	Φ^{298K} (± 0.05)
TDSiNPs	*a*	3.4	3.4	0.67	0.5 (76)	0.2 (air: 0.005)
	b	3.15	3.1	0.53	3.15 (19)	0.19
PA-TDSiNPs	*c*	3.0	2.95	0.49	1.6 ± 0.1 (70)	0.16
	d	2.7	2.7	0.50	4.3 ± 0.1 (27)	0.2 (air: 0.14)
SiO$_x$-TDSiNPs	***i***	**3.0**	**3.0**	**0.45**	**1.27 ± 0.05**	**0.13**
	j	**3.4**	**3.5**	**0.59**	**-----**	**----**
PA-SiO$_x$-TDSiNPs	***k***	**3.0**	**3.0**	**0.45**	**1.30 ± 0.05**	**0.36**
	l	**3.4**	**3.4**	**0.60**	**----**	**0.3**
MASiO$_x$-TD′SiNPs[b]	***m***	**3.0**	**2.85**	**0.43**	**1.19 ± 0.05**	**0.76 (air:0.44)**
	n	**3.3**	**3.4**	**0.62**	**----**	**-----**
BUSiNPs	*e*	3.1	3.1	0.55	1.5 ± 0.5	-----
	f	2.8	2.8	0.45	7 ± 0.5	-----
PA-BUSiNPs	*g*	3.15	3.1	0.56	1.2 ± 0.3 (80)	0.25
	h	2.65	2.75	0.63	4.5 ± 0.1 (20)	0.34
SiO$_x$-BUSiNPs[b]		**3.0**	**2.85**	**0.45**	**1.6 ± 0.05**	**------**
PASiO$_x$-BUSiNPs	***o***	**2.95**	**2.85**	**0.44**	**1.1 ± 0.1**	**0.1**

[a]Depicted Φ are lower limit values, see text. Values in bold stand for surface oxidized particles. [b]Data from ref 16.

368 expected k_r values fall within $(1-60) \times 10^7$ s^{-1}. However, as it
369 will be discussed later, Φ values also consider the absorbance of
370 dark particles, and therefore calculated k_r values are lower limit
371 estimations.

372 The mean energy gap (EG) of charge carriers in silicon
373 nanoparticles may be evaluated using the excitation spectrum

(PLE) threshold of each individual species.[16,34] Calculated EG 374
values for each emitting species, shown in Table 2, coincide 375
with the PL maxima, strongly supporting that the excitation and 376
the emission mainly originate from the same transition. 377
Corresponding Stokes shifts (SS), also shown in Table 2, 378
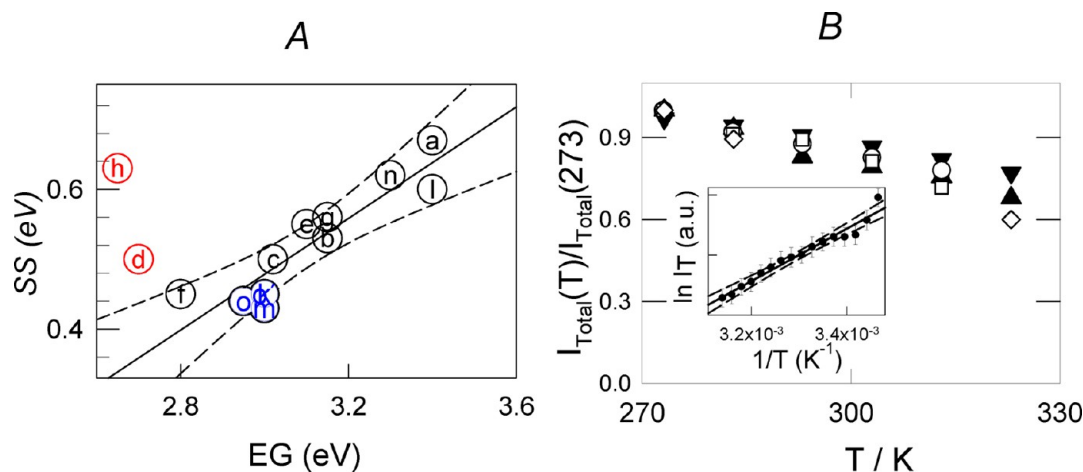vary within the 0.4–0.6 eV range. 379

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Figure 4. (A) Stokes distance dependence with the energy gap. The letters inside the symbols stand for the different emitting species. (B) Ratio of the total emission intensity at a given temperature to that observed at 273 K ($I_T(T)/I_T(273)$) for argon-saturated toluene suspensions of species *a* (▼), *b* (▲), *c* (◇), *k* (○), and *m* (□) vs T. Inset: ln I_T vs T^{-1} for argon-saturated toluene suspensions of species *k*.

The PL of TDSiNPs suspensions shows the contribution of two emitting species (*a* and *b* in Figure 3A,B). The observed wide PL bands (fwhm of about $\Delta E \approx 400$ meV) extending over the 350−450 nm range and decay lifetimes in the nanosecond regime are in agreement with those reported for 1 nm size SiNPs.[8,9,35] On the assumption that the observed luminescence is controlled by quantum confinement, particle sizes of 1.3−1.6 nm are estimated from theoretical correlations of the EG with size reported for H-terminated crystalline SiNPs,[7] in line with experimentally determined values (vide supra). The PL of PA-TDSiNPs also shows two contributing species (*c* and *d* in Figure 3A,B) with red-shifted PL compared to those of *a* and *b*. Since PA-TDSiNPs and TDSiNPs mainly differ in their surface composition ($\equiv$Si—CH$_2$—CH$_2$—CH$_2$−NH$_2$ and $\equiv$Si-NH—CH$_2$—CH$=$CH$_2$ surface groups in place of $\equiv$Si—H), the observed differences in their excitation−emission matrix may only be attributed to the strong dependence of the PL on the surface chemical composition. Recent theoretical and experimental studies indicate that coverage of SiNPs with alkyl-groups (Si—C) results in a minimal change in PL spectra. Since SiNPs with Si—N terminations exhibit significant PL shifts compared to H-passivated SiNPs,[27,36] the observed differences may be attributed to Si—N surface bonds.

On the other hand, the PL and PLE spectra of amorphous BUSiNPs (*e* and *f* in Figure 3C,D) are shifted to the red with respect to those of crystalline TDSiNPs, as expected from theoretical calculations for amorphous[37] and crystalline[7] particles of similar size. It has been recently reported[27] that H-passivated SiNPs exposed to the surfactant tetraoctylammonium bromide yield PL spectra in the 400 and 450 nm range as a consequence of the formation of Si—N surface bonds. Since no evidence on the formation of Si—N environments was obtained from XPS and FTIR data of BUSiNPs,[16] tetraoctylammonium bromide chemisorption to the particle surface is of no significance under our synthesis conditions. In fact, the PL of species *e* and *f* is coincident to that reported for 1.5 nm-size amorphous SiNPs obtained from the electrochemical etching of amorphous Si films.[38] The excitation−emission matrix of propylamine-surface derivatized BUSiNPs shows the contribution of two emitting species (*g* and *h* in Figure 3C,D) with similar PL and PLE spectra to those of species *e* and *f*. Thus, Si—C and Si—N environments confining BUSiNPs appears

not to significantly affect the overall density of states within the capture volume of spatially confined carriers leading to PL in amorphous SiNPs.[39]

Surface oxidized particles of diverse origin (TD and BU synthesis) and with different capping groups, mainly SiO$_x$-TDSiNPs, PASiO$_x$-TDSiNPs, SiO$_x$-TD′SiNPs, MASiO$_x$-TD′SiNPs, SiO$_x$-BUSiNPs, and PASiO$_x$-BUSiNPs, show a similar excitation−emission matrix (see Figure 3E,F) and PL decay times τ (see Table 2). These observations strongly indicate that the excitation−emission matrix shape and PL decay times of oxidized SiNPs < 3 nm in size with surface $\equiv$Si—O—Si$\equiv$ structures does not depend on the particle crystalline structure, neither on size nor on further organic capping. The presence of surface $\equiv$Si—O—Si$\equiv$ structures leads to an EG reduction in the order of 0.2−0.4 eV with respect to the respective H-passivated TDSiNPs. These observations are in line with theoretical studies predicting that the addition of a single oxygen atom to the surface of 1 nm size H-passivated Si clusters lowers the optical absorption gap by fractions of eV when attached in a bridged configuration (Si—O—Si) but larger impacts on the EG (units of eV) are expected when attached in a double-bonded ($>$Si$=$O) configuration.[8,10]

The excitation and emission spectrum of the oxidized particles shows peaking with an average peak separation of 140−180 meV, on the order of the splitting due to zero-phonon line and phonon side bands reported in single molecule experiments of 5 and 6 nm size oxidized crystalline SiNPs embedded in a matrix.[40−43] Energy spacing close to 150 meV may be attributed to Si—O—Si longitudinal optical (LO) and transverse optical (TO) phonons at the interface.[41] Assuming an electron−phonon interaction, the higher energy of the zero phonon line observed herein ($>$3.0 eV) might correlate with the smaller silicon core size of the particles in a quantum confinement regime but seems not to support a strong localization of both the electron and the hole at the interface as suggested by the observed independence on particle size and crystallinity.[42] Since effects as the compressive strain exerted by the matrix on embedded nanocrystals may also influence Si-NPs PL,[44] any discussion on the origin of the observed differences requires the comparison with single molecule experiments with the ultrasmall blue emitting SiNPs.

Overall, the obtained results strongly indicate that for small structures where the number of surface atoms is a significant fraction of the total number of atoms (see S.I. Surface Si atoms estimation (Supporting Information)), different adsorbates and atomic arrangements lead to different surface electronic states which influence the SiNPs PL. SS variations of crystalline SiNPs have been largely attributed to energy levels below the conduction band minimum leading to self-trapped excitons localized in the surface. Such deep levels were reported to be introduced by surface states, dangling bonds, and electron–phonon interactions.[45] Interestingly, Figure 4A shows an acceptable correlation between the SS and the EG for the different emitting species. These observations and the fact that the same transition is involved in the absorption and emission processes (vide supra) strongly support the contribution of surface states to the energy levels in the gap. It should be noted that species *d* and *h* contributing to PA-TDSiNPs and PA-BUSiNPs, respectively, are an exception. Chemisorbed N may introduce additional deep levels below the conduction band minimum not affecting the EG. Unfortunately, determination of the precise nature of the contributing species is not possible, and the data is usually open to various interpretations.

Larger SiNPs (>2 nm size) explored in the literature showing green to near-IR PL originated in net quantum confinement effects[44,46,47] are also reported to show important modifications in the PL upon changes in their surface chemistry.[48−50] Therefore, the participation of surface states in SiNPs PL exceeds our observations for small silicon nanoparticles with large surface to volume ratios.

Temperature Dependence of the Emission Quantum Yield. The overall luminescence quantum yield of argon-saturated toluene suspensions of the particles diminish with increasing temperature, as shown for MASiO$_x$-TD′SiNPs in S.I. Table 1 (Supporting Information). The Φ values are fully recovered upon restoring the temperature to the initial values, thus supporting reversible temperature-dependent quenching processes. Plots of $\ln(\Phi^{-1} - 1)$ vs T^{-1} show curvature, as expected if two or more deactivation processes with significant activation energies are present. In fact, Φ values cannot separate the individual contributions of all the emitting species. Therefore, to understand the temperature effect on PL, a bilinear analysis of the excitation–emission obtained at each temperature in the range from 273 to 330 K was performed for all the particle samples. Such analysis indicates that the intensity but neither the EG nor the SS of each contributing species depends on temperature. Figure 4B shows the decrease of I_T/I_{273} with temperature for several of the contributing species. The ratio of total intensity at a given temperature T with respect to that observed at 273 K (I_T/I_{273}) was used to compensate for all non-temperature-dependent parameters affecting I_T as those of the equipment response and particle concentration. Except for those of species *d*, I_T/I_{273K} values fall within the same curve, therefore suggesting that the thermally activated quenching does not show a correlation with particle structure and surface chemistry. Interestingly, τ values showed no variation, within the experimental error, in the 200–330 K temperature range.

Calcott's model assuming a thermal equilibrium between two excitonic levels of different multiplicity and radiative constant, split by energy (ΔE), is widely used in the literature to interpret the dependence with temperature of the radiative rate constant k_r.[13,14,51−53] At very low temperatures, almost all of the excitons are trapped in the lower state, whereas at higher temperatures, the excitons are in their upper state which is a dominant luminescent center. Reported ΔE values are on the order of 3 meV.[13,52] Considering that k_r is expected not to vary with temperature for $T > 200$ K, decreasing PL intensities with increasing temperature have been attributed to the dominating role of nonradiative recombination.[13,54] Two main models were postulated to account for nonradiative thermally activated processes: a normal Arrhenius behavior[13,55] and the ionization of the excited state via tunneling through a potential barrier with a Berthelot-type T-dependence.[54,56−58] However, considering that our results show that $(k_r + \sum_i k_{ni}) = \tau^{-1}$ remains constant with temperature, k_r being itself T-independent, competing nonradiative processes $\sum_i k_{ni}$ originated in the excitonic level are also expected to be temperature independent. Therefore, the involvement of any of the described mechanisms cannot describe the observed T-dependence of the PL intensity.

A static, reversible, thermal quenching of the particles is strongly suggested. Increasing temperatures lead to an increased population of nonemitting particles which might be reverted upon a decrease in temperature. Moreover, since no correlation between the quenching rate and the particle surface chemistry is observed, it is expected that the silicon structure is involved in leading to dark (nonemitting) clusters. These observations are in agreement with those by Diener and co-workers[59] indicating that IR excitation of porous silicon results in a drop in the PL intensity but only with a slight change in the PL lifetime. The authors suggested that IR excitation creates additional dangling bonds or opens additional nonradiative channels leading to an increase of the fraction number of dark crystallites.

Plots of $\ln I_T$ vs T^{-1} (see inset in Figure 4B) yield straight lines from which the slopes of the enthalpy change $\Delta H = +(15 \pm 1)$ meV of the thermal process leading to the dark particle conformation may be obtained. The observed energy difference between a bright and a dark particle is too small to be assigned to the generation of dangling bonds and are thus suggested to be due to small changes in the structure conformation.

Literature reports on the presence of a dark fraction of particles in quantum dots have been related to blinking dynamics.[60] Reported single molecule level PL studies with Si nanocrystals revealed PL blinking with power-law statistics and "on" and "off" times in the millisecond–second range.[61] Strong disorder effects were reported to provoke the reversible bleaching of SiNPs on a time scale of hours in the dark.[61] A photogeneration mechanism of the dark particles would imply a reduction in the intensity of steady-state PL with the irradiation time. Since no changes in the steady state PL intensity were observed with the illumination time for the particles herein synthesized, dark SiNPs might be assumed in equilibrium with bright particles before illumination. Scheme 3 illustrates the thermal process between dark and bright particles leading to an increased fraction of dark particles as temperature increases. If power-law statistics of blinking also holds for our SiNPs, then it could arise from a multiplicity of dark (or bright) conformations with different energy barriers.

■ CONCLUSION

Important challenges in the use of SiNPs as sensors and photosensitizers are the following: their PL dependence on size, crystallinity, and surface structure and chemistry. Also, the low stability of H-terminated SiNPs leads to unwanted surface-oxidized particles unless the synthesis process and surface

Scheme 3. Illustration of the Thermal Process Involving Dark and Bright SiNPs (not scaled)[a]
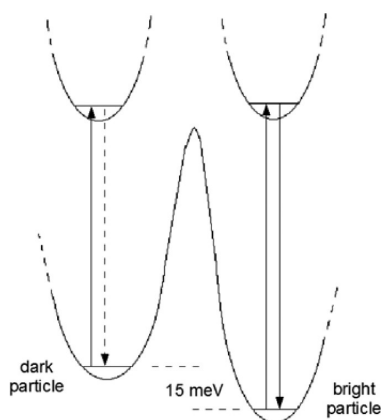
[a]Arrows stand for radiative (solid) and non-radiative (dashed) processes.

derivatization are performed under strict oxygen-free conditions. Our results show that the PL properties of mild surface oxidized 1−2 nm sized particles mainly showing Si—O—Si bridging bonds are independent of the synthesis procedure (BU or TD) and on the further derivatization of the particles surface with organic molecules with different terminal groups. This observation has important consequences for the particles' technological uses as the proper synthesis route may be selected on costs and availability of reactants and equipment. Moreover, even though the main synthetic procedure requires minimization of the dissolved O_2, further silanization procedures for organic capping may be conducted under an air atmosphere at room temperature. As a result, biocompatible molecules may be attached to the surface without changes in the PL. The latter properties are important for specific targeting and opsonin association prevention in sensing applications.

■ ASSOCIATED CONTENT

Ⓢ Supporting Information
TEM, XPS, and FTIR data. This material is available free of charge via the Internet at http://pubs.acs.org.

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Notes
The authors declare no competing financial interest.

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