

Theoretical Study of the Electronic Spectrum of Disulfur Monoxide

Carlos J. Cobos and Adela E. Croce

The Research Institute of Theoretical and Applied Physical Chemistry, Department of Chemistry, Faculty of Exact Sciences, National University of La Plata, Casilla de Correo 16, Sucursal 4, (1900) La Plata, Argentina

Reprint requests to A. E. C.; E-mail: acroce@inifta.unlp.edu.ar

Z. Naturforsch. **69a**, 215–219 (2014) / DOI: 10.5560/ZNA.2014-0009

Received October 29, 2013 / revised February 5, 2014 / published online April 2, 2014

The near ultraviolet-visible absorption spectrum of disulfur monoxide (S_2O) has been theoretically studied by using the time-dependent density functional theory (TD-DFT) and the equation of motion coupled-cluster singles and doubles approach (EOM-CCSD) combined with the AUG-cc-PVQZ basis set. From this, analytical expressions for the absorption coefficient over the 250–340 nm range are reported for the first time. The computed molecular structure and the vibrational frequencies for the ground and third electronically excited state S_2O (C^1A'), responsible of the observed spectrum, are compared with available data.

Key words: S_2O ; Absorption Spectra; TD-DFT; EOM-CCSD.

1. Introduction

In addition to the well identified sulfur dioxide (SO_2) molecule, the presence of disulfur monoxide (S_2O) has been proposed to explain some relevant spectral features of the surface of Io, Jupiter's moon. Modelling calculations suggest the formation of up to 6% of S_2O from equilibrated SO_2/S_2 mixtures near the volcanoes of Io [1]. This species also participates in the Venus sulfur chemistry, which determines the composition of the middle atmosphere [2]. The sulfur oxide obtained by streaming SO_2 at low pressure through a high voltage electric discharge [3] is at present known as disulfur monoxide S_2O [4, 5]. The electronic absorption spectrum [6–10] has been investigated in detail. The S_2O exhibits two absorption regions, one extending from 250 to 340 nm, and the other from 190 to 230 nm. In particular, the first region presents an intense and richly-structured absorption band picket at about 295 nm. Although vibrational analysis studies of this band have been reported [6–10], surprisingly no absorption coefficient data are available. Due to the fact that this information is relevant for future photochemical, kinetic, and modelling studies of planetary importance, a theoretical study on the basis of the complete near-UV spectrum of S_2O measured by Phillips et al. [8] has been undertaken.

2. Computational Methods

Electronic excitation energies of valence excited states which can be described by combinations of single one electron transitions can be reasonably well described by the time-dependent density functional theory (TD-DFT) theory. However, due to the fact that a mathematical expression for the real exchange-correlation functional in DFT is unknown [11], a large number of specific and general approaches have been proposed. For this reason, different formulations are usually employed and the resulting data are normally confronted with experiments. In this study the following hybrid functionals were employed: B3LYP, O3LYP, X3LYP, B1LYP, mPW1LYP, B3PW91, B97-2, B98, PBE1PBE, mPW1PW91, BMK, M06-2X, tHCTHhyb, wB97XD, CAM-B3LYP, and HSEh1PBE. In addition, the excitation energies were computed using the state-of-the-art equation of motion coupled-cluster singles and doubles method (EOM-CCSD). For all models the large AUG-cc-pVQZ basis set was used [12]. The present EOM-CCSD calculations are based on B3LYP optimized geometries. All mentioned methods are implemented in the employed Gaussian suite of programs [13].

3. Absorption Spectrum

The S_2O absorption spectrum of S_2O trapped in a xenon matrix at 20 K (up to three units of absorbance

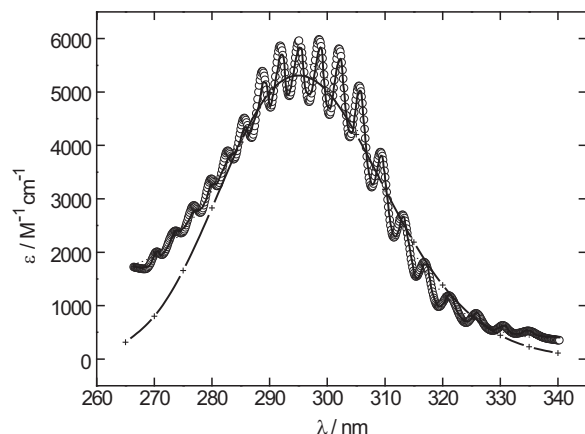


Fig. 1. Comparison between the experimental and computed ultraviolet-visible spectrum of S_2O . (–): Experiments from [6]; (–+–): Present quantum-mechanical calculations (see text); (o): (2).

scale) is depicted in Figure 1 [8]. The spectrum consists in a single progression of vibrational bands (S–S stretching mode) very similar to the observed in the gaseous phase and at room temperature [6, 7]. It is interesting to note that a recent low temperature matrix spectrum, only limited to the 320–340 nm range, shows an identical structure for the first four vibrational levels of S_2O [14]. Our calculations show that to cover the near-UV region of the experimental spectrum, only the third singlet-singlet valence electronic transition of S_2O is required. In fact, the first two EOM-CCSD electronic transitions, as indicated by the very low oscillator strengths f , exhibit negligible intensity and, above the third, the more important transitions are the sixth and the eighth located at 178 nm ($f = 0.0089$) and 167 nm ($f = 0.0717$), respectively.

The vertical excitation energies E_m , the associated wavelengths λ_m , the f values, the structural parameters, and harmonic vibrational frequencies for both, the ground and the third electronically excited state are listed in Table 1. Additional TD-DFT calculations carried out with different basis sets lead to essentially similar results. By instance, the alternative use of the 6-311+G(3d2f) Pople's basis set [15], of the AUG-cc-pVQZ basis set with the inclusion of tight d-functions or 2df-functions, or even of the larger AUG-cc-pV5Z basis set does not affect λ_m and f significantly. In fact, the resulting values obtained with the B3LYP functional are: $\lambda_m = 284$ nm and $f = 0.0646$ (6-311+G(3d2f)); $\lambda_m = 284$ nm and $f = 0.0644$ (AUG-cc-pV(Q+d)Z); $\lambda_m = 284$ nm and $f = 0.0642$ (AUG-cc-pV(Q+2df)Z);

and $\lambda_m = 288$ nm and $f = 0.0640$ (AUG-cc-pV5Z). As Table 1 shows, these values are in very good agreement with those obtained with the AUG-cc-pVQZ basis set.

Despite the different generation of exchange and correlation functionals employed, similar results were found and, thus, no approach can be definitely attributed superior to the rest. Therefore, for simplicity, no individual theoretical method is discussed. All data were averaged to facilitate the comparison with available experimental and theoretical data. In this way, the TD-DFT average values of $E_m = 4.38 \pm 0.06$ eV ($\lambda_m = 283 \pm 4$ nm) and $f = 0.0697 \pm 0.005$ were derived. As Table 1 shows, the EOM-CCSD value for the band localization, 4.44 eV ($\lambda_m = 279$ nm), is within the above given deviation. However, the oscillator strength is $\approx 25\%$ larger than those provided by the TD-DFT methods. The theoretical predictions, when compared with the experimental value of about 4.2 eV (≈ 295 nm), lead to a center of band position shifted in about $+0.2$ eV. A similar behaviour has been observed for the three absorption bands of the CF_3OSO_3 radical [16]. It should be noted that typical deviations of 0.3–0.4 eV (TD-DFT) [17, 18] and 0.2–0.3 eV (EOM-CCSD) [18] have been reported for a large number of small- to medium-sized molecules. On the other hand, all computed f values of Table 1 lie in between the notably discrepant values of 0.17 [19] and ≈ 0.026 [20] derived from measured radiative lifetimes of the C^1A' state.

The calculations indicate that the largest spin coefficient for the third excited state is $c_{18 \rightarrow 21} = 0.670 \pm 0.001$ (average of all employed methods), and corresponds to the excitation from the molecular orbital 18 to the molecular orbital 21 ($\pi \rightarrow \pi^*$ transition located along the S–S bond). The resulting contribution percentage of this configuration is given by $100 \times 2 \times c_{18 \rightarrow 21}^2 = 90\%$, where the factor 2 accounts for two contributions in a properly spin-adapted eigenfunction.

To a best comparison between experimental and theoretical results, the decadic absorption coefficient ε (in units $M^{-1} \text{ cm}^{-1}$) over the relevant energy range was calculated as [16]

$$\varepsilon \cong 1.15 \cdot 10^4 \left(\frac{f}{\sigma} \right) \exp \left(-\frac{1}{2} \left(\frac{E - E_m}{\sigma} \right)^2 \right). \quad (1)$$

Here the band form is represented by a Gaussian function $\varepsilon = \varepsilon_m \exp(-\frac{1}{2}(\frac{E-E_m}{\sigma})^2)$, with a full width of the

band at $\frac{1}{e}$ height of σ , and an energy maximum of $E_m = \hbar c/\lambda_m$, where \hbar is Planck's constant and c the velocity of light in a vacuum. As usual, the σ parameter was used to match the theoretical and experimental spectra.

To compare the theoretical with the experimental data, the transformation of the absorbances (measurements performed up to three units of absorbance scale [6]) to absorption coefficients is required. However, because neither the optical pathway l as defined by the thickness of the S_2O deposited in the cryogenic matrix nor the S_2O concentration c are known, the absorbances were appropriately scaled to transform to absorption coefficients [16]. In fact, an scaling factor of $\frac{1}{lc} = 7634 \text{ M}^{-1} \text{ cm}^{-1}$ has been recently employed to reproduce over the 220–540 nm range, the spectrum of

the CF_3OSO_3 radical deposited in a neon matrix at 6 K [16]. For the present case, we have modelled the spectrum employing the f value predicted by the EOM-CCSD calculations and shifted the computed average λ_m of 283 nm in only -12 nm to superpose the maxima values of experimental and theoretical curves. In the absence of a reliable theoretical foundation for the vibrational broadening, the bandwidth $\sigma = 0.20$ eV and the scaling factor $1950 \text{ M}^{-1} \text{ cm}^{-1}$ were employed to account for the spectrum shape. It should be noted that this σ value is almost equal to the found for the more complicated three-band spectrum of the CF_3OSO_3 radical of 0.21 eV [16].

As Figure 1 shows, neglecting the vibronic structure, the resulting theoretical curve reproduces acceptably well the global contour of the spectrum above

Table 1. Vertical excitation energies, wavelengths, oscillator strengths, S–S and S–O bond lengths, bond angles, and the corresponding harmonic vibrational frequencies (see text). For all calculations, the basis set AUG-cc-pVQZ was used. Each first line gives the values for the electronically excited state, each second line those for the ground electronic state.

Level of theory	E_m (eV)	λ_m (nm)	f	Bond lengths (Å) and angles (°)	Frequencies (cm ⁻¹)
B3LYP	4.31	288	0.0656	2.176, 1.486, 109.1 1.901, 1.465, 117.8	384, 1080, 243 682, 1182, 381
O3LYP	4.38	283	0.0595	2.145, 1.490, 109.2 1.895, 1.466, 118.0	400, 1065, 251 690, 1186, 381
X3LYP	4.32	287	0.0666	2.171, 1.484, 108.9 1.899, 1.463, 117.7	388, 1087, 245 686, 1190, 383
B1LYP	4.32	287	0.0692	2.173, 1.482, 109.0 1.898, 1.462, 117.7	387, 1096, 246 690, 1196, 385
mPW1LYP	4.32	287	0.0685	2.171, 1.482, 108.9 1.897, 1.461, 117.6	388, 1096, 246 690, 1198, 385
B3PW91	4.41	281	0.0673	2.140, 1.485, 108.9 1.891, 1.460, 117.7	409, 1086, 255 701, 1205, 385
B97-2	4.46	278	0.0687	2.132, 1.481, 109.0 1.885, 1.457, 117.8	409, 1096, 258 712, 1215, 390
B98	4.32	287	0.0670	2.171, 1.481, 109.0 1.904, 1.460, 117.7	403, 1107, 250 677, 1206, 382
PBE1PBE	4.47	277	0.0702	2.124, 1.480, 108.6 1.885, 1.456, 117.7	421, 1107, 261 718, 1227, 391
mPW1PW91	4.46	278	0.0705	2.128, 1.479, 108.8 1.885, 1.455, 117.7	418, 1107, 260 715, 1226, 391
BMK	4.34	286	0.0742	2.149, 1.469, 108.7 1.913, 1.448, 117.5	455, 1153, 267 700, 1242, 384
M06-2X	4.40	282	0.0790	2.122, 1.472, 108.4 1.877, 1.450, 117.4	421, 1160, 263 730, 1263, 394
tHCTHhyb	4.33	286	0.0630	2.162, 1.486, 108.9 1.903, 1.463, 117.8	400, 1079, 249 677, 1186, 378
wB97XD	4.45	279	0.0788	2.125, 1.473, 108.6 1.879, 1.450, 117.5	430, 1132, 267 725, 1259, 397
CAM-B3LYP	4.40	282	0.0781	2.128, 1.474, 1.805 1.878, 1.452, 117.5	417, 1129, 261 726, 1240, 396
HSEh1PBE	4.46	278	0.0695	2.128, 1.480, 108.7 1.886, 1.457, 117.7	417, 1101, 259 714, 1221, 389
EOM-CCSD	4.44	279	0.0935	–	–

≈ 285 nm. As above mentioned, no relevant electronic transitions are present up to about 6.9 eV (≈ 180 nm), therefore, the departure from the simple Gaussian shape could be attributed to the well established predissociation within the electronically-excited manifold, leading to ground state S and SO species [21].

To practical purposes (i. e., for photochemical and kinetics studies of second-order reactions) realistic values of ε vs. λ are highly desirable. In the absence of direct measurements of the S₂O absorption coefficient, and on the basis of the above analysis, the spectrum can be accurately fitted by a sum of single Gaussian functions. The resulting function, which accounts for the vibrational structure of the spectrum, is the following:

$$\varepsilon = -4.599 \cdot 10^3 + 1.684 \cdot 10^{-1} \nu + 7.979 \cdot 10^{-1} \sum_{i=0}^{18} \frac{A_i}{\omega_i} \exp\left(-2\left(\frac{\nu - \nu_i}{\omega_i}\right)^2\right), \quad (2)$$

where $\nu = \frac{1}{\lambda}$. The curve depicted in Figure 1 has been obtained using the A_i , ω_i , and ν_i specific coefficients for the i th vibrational state listed in Table 2.

4. Molecular Structure and Harmonic Vibrational Frequencies

The extensive band system suggests a substantial geometry change after the electronic excitation to the third

electronic state of S₂O. In fact, the light absorption by the S–S double bond promotes one π valence electron into a virtual antibonding π^* orbital. As a consequence, the concomitant S–S elongation and reduction of the S–S–O angle are expected. Actually, from the resulting DFT structural data listed in Table 1, the following average values were obtained: $r_{S-O} = 1.480 \pm 0.006$ Å, $r_{S-S} = 2.146 \pm 0.021$ Å, and $\angle S-S-O = 108.8 \pm 0.2^\circ$. These values compare very well with the experimental values of 1.50 Å (assumed value), 2.14 Å, and $\approx 109^\circ$ [10], and 1.483 Å, 2.168 Å, and 106 or 110° [21]. Similar calculations performed for the ground state (see Tab. 1) lead to $r_{S-O} = 1.458 \pm 0.006$ Å, $r_{S-S} = 1.892 \pm 0.011$ Å, and $\angle S-S-O = 117.7 \pm 0.1^\circ$, in very good agreement, respectively, with the measured data of 1.4594 Å, 1.8845 Å, and 118.08° [5, 22].

As Table 1 shows, after excitation, the S–S bond weakening yields a marked reduction in the frequency values. The resulting average values for the S–S stretching, S–O stretching, and bending modes are: 409 ± 19 , 1105 ± 26 , and 255 ± 8 cm⁻¹ (excited state) and 702 ± 18 , 1215 ± 26 , and 387 ± 6 cm⁻¹ (ground state). For both electronic states, the S–S stretching and bending degrees of freedom agree reasonably well with the infrared values of 405 [19, 23] and 415.2 cm⁻¹ [20], and 252 [19] and 253 cm⁻¹ [23] (excited state) and 679 and 388 cm⁻¹ (ground state) [5, 18, 20]. However, our computed values for the S–O stretching overestimates the experimental values of 1030 [19] and 1032 cm⁻¹ [23] (excited state), and 1165 cm⁻¹ (ground state) [7, 22, 24].

Table 2. State-selected coefficients for (2).

Vibrational state	A_i	ω_i (cm ⁻¹)	ν_i (cm ⁻¹)
0	$2.587 \cdot 10^4$	$2.146 \cdot 10^2$	$2.985 \cdot 10^4$
1	$2.307 \cdot 10^4$	$1.385 \cdot 10^2$	$3.026 \cdot 10^4$
2	$6.185 \cdot 10^4$	$1.741 \cdot 10^2$	$3.069 \cdot 10^4$
3	$1.673 \cdot 10^5$	$2.471 \cdot 10^2$	$3.114 \cdot 10^4$
4	$3.007 \cdot 10^5$	$2.241 \cdot 10^2$	$3.155 \cdot 10^4$
5	$6.025 \cdot 10^5$	$2.559 \cdot 10^2$	$3.194 \cdot 10^4$
6	$9.758 \cdot 10^5$	$2.645 \cdot 10^2$	$3.233 \cdot 10^4$
7	$1.413 \cdot 10^6$	$2.751 \cdot 10^2$	$3.273 \cdot 10^4$
8	$1.510 \cdot 10^6$	$2.619 \cdot 10^2$	$3.310 \cdot 10^4$
9	$1.700 \cdot 10^6$	$2.870 \cdot 10^2$	$3.348 \cdot 10^4$
10	$1.790 \cdot 10^6$	$3.058 \cdot 10^2$	$3.389 \cdot 10^4$
11	$1.496 \cdot 10^6$	$2.746 \cdot 10^2$	$3.427 \cdot 10^4$
12	$1.387 \cdot 10^6$	$2.833 \cdot 10^2$	$3.463 \cdot 10^4$
13	$1.145 \cdot 10^6$	$3.024 \cdot 10^2$	$3.502 \cdot 10^4$
14	$8.403 \cdot 10^5$	$2.903 \cdot 10^2$	$3.539 \cdot 10^4$
15	$6.223 \cdot 10^5$	$2.804 \cdot 10^2$	$3.574 \cdot 10^4$
16	$4.992 \cdot 10^5$	$3.089 \cdot 10^2$	$3.612 \cdot 10^4$
17	$3.806 \cdot 10^5$	$3.728 \cdot 10^2$	$3.655 \cdot 10^4$
18	$8.065 \cdot 10^4$	$2.006 \cdot 10^2$	$3.699 \cdot 10^4$

5. Conclusion

This study allows to reproduce the observed features of the S₂O spectrum between 270 and 340 nm and to transform the absorbance data of [6] on an absolute basis, useful for future kinetics and planetary simulations. Besides, the obtained structural and infrared spectroscopic properties of the S₂O (C¹A') are quite consistent with the experimental ones, supporting the presented theoretical analysis of the spectrum.

Acknowledgements

This work was supported by the National University of La Plata and the Consejo Nacional de Investigaciones Científicas y Técnicas (CONICET).

- [1] M. Yu. Zolotov and B. Fegley, Jr., *Icarus* **133**, 293 (1998).
- [2] X. Zhang, M. C. Liang, F. P. Mills, D. A. Belyaev, and Y. L. Yung, *Icarus* **217**, 714 (2012).
- [3] P. W. Schenk, *Z. Anorg. Allg. Chem.* **211**, 150 (1933).
- [4] D. J. Meschi and R. J. Myers, *J. Am. Chem. Soc.* **78**, 6220 (1956).
- [5] D. J. Meschi and R. J. Myers, *J. Mol. Spectrosc.* **3**, 405 (1959).
- [6] H. Cordes, *Z. Physik* **105**, 251 (1937).
- [7] A. V. Jones, *J. Chem. Phys.* **18**, 1263 (1950).
- [8] L. F. Phillips, J. J. Smith, and B. Meyer, *J. Mol. Spectrosc.* **29**, 230 (1969).
- [9] G. Lakshminarayana, *J. Mol. Spectrosc.* **55**, 141 (1975).
- [10] K.-E. J. Hallin, A. J. Merer, and D. J. Milton, *Can. J. Phys.* **55**, 1858 (1977).
- [11] P. Hohenberg and W. Kohn, *Phys. Rev. B* **136**, 864 (1964).
- [12] D. E. Wong and T. H. Dunning, *J. Chem. Phys.* **98**, 1358 (1993); and references therein.
- [13] M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A. Petersson, H. Nakatsuji, M. Caricato, X. Li, H. P. Hratchian, A. F. Izmaylov, J. Bloino, G. Zheng, J. L. Sonnenberg, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, J. A. Montgomery, Jr., J. E. Peralta, F. Ogliaro, M. Bearpark, J. J. Heyd, E. Brothers, K. N. Kudin, V. N. Staroverov, R. Kobayashi, J. Normand, K. Raghavachari, A. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, N. Rega, J. M. Millam, M. Klene, J. E. Knox, J. B. Cross, V. Bakken, C. Adamo, J. Jaramillo, R. Gomperts, R. E. Stratmann, O. Yazyev, A. J. Austin, R. Cammi, C. Pomelli, J. W. Ochterski, R. L. Martin, K. Morokuma, V. G. Zakrzewski, G. A. Voth, P. Salvador, J. J. Dannenberg, S. Dapprich, A. D. Daniels, O. Farkas, J. B. Foresman, J. V. Ortiz, J. Cioslowski, and D. J. Fox, Gaussian 09, Revision A.02, Gaussian Inc., Wallingford CT, 2009.
- [14] W.-J. Lo, Y.-J. Wu, and Y.-P. Lee, *J. Phys. Chem. A* **107**, 6944 (2003).
- [15] M. J. Frisch, J. A. Pople, and J. S. Binkley, *J. Chem. Phys.* **80**, 3265 (1984); and references therein.
- [16] C. J. Cobos and A. E. Croce, *Z. Naturforsch.* **65a**, 720 (2010).
- [17] S. S. Leang, F. Zahariev, and M. S. Gordon, *J. Chem. Phys.* **136**, 104101 (2012).
- [18] M. Caricato, G. W. Trucks, M. J. Frisch, and K. B. Wiberg, *J. Chem. Theory Comput.* **6**, 370 (2010).
- [19] K. Tsukiyama, D. Kobayashi, K. Obi, and I. Tanaka, *Chem. Phys.* **84**, 337 (1984).
- [20] Q. Zhang, P. Dupré, B. Grzybowski, and P. H. Vaccaro, *J. Chem. Phys.* **103**, 37 (1995).
- [21] T. Müller, P. H. Vaccaro, F. Pérez-Bernalá, and F. Iachello, *J. Chem. Phys.* **111**, 5038 (1999).
- [22] J. Lindenmayer, H. D. Rudolph, and H. Jones, *J. Mol. Spectrosc.* **119**, 56 (1986); and references therein.
- [23] D. J. Clouthier and M. L. Rutherford, *Chem. Phys.* **127**, 189 (1988).
- [24] U. Blukis and R. J. Myers, *J. Phys. Chem.* **69**, 1154 (1965).