

Thermal reversion of spirooxazine in ionic liquids containing the [NTf₂]⁻ anion

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In the above paper, Table 1 inadvertently referenced the incorrect literature. The references should appear as shown in the table below (according to the numbering in the published paper). The literature value for α in ethanol is corrected to 0.86 from 0.83 and the literature value for β in ethanol is corrected to 0.75 from 0.77.

Solvent	Et(30)/kcal mol ⁻¹	α	β	π^*	SO			
					MC λ_{\max}	$k \times 10^{-2}/s^{-1}$	SD	τ/s
Methanol	55.4 (55.4) ⁸	1.06 (1.05) ⁴¹	0.62 (0.61) ⁴¹	0.71 (0.73) ⁴¹	640	3.2	±0.008	31.25
Ethanol	52.1 (51.9) ⁸	0.90 (0.86) ⁴²	0.72 (0.75) ⁴²	0.63 (0.54) ⁴²	642	3.8 (2.0) ²⁰	±0.009	26.32
Acetonitrile	46.4 (45.6) ⁸	0.42 (0.35) ⁴¹	0.37 (0.37) ⁴¹	0.79 (0.79) ⁴¹	642	5.0 (5.2) ²⁰	±0.006	20
Acetone	42.5 (42.2) ⁸	0.25 (0.20) ⁴¹	0.57 (0.54) ⁴¹	0.67 (0.70) ⁴¹	642	5.1 (5.4) ²⁰	±0.011	19.61
[bmIm][NTf ₂]	52.4 (51.5) ⁴³	0.72 (0.69) ⁴¹	0.24 (0.25) ⁴¹	0.90 (0.97) ⁴¹	642	2	±0.005	50
[em ₂ Im][NTf ₂]	50	0.42	0.1	1.02	640	2.3	±0.011	43.48
[bmPy][NTf ₂]	49.6 (50.2) ¹⁴	0.57 (0.43) ¹³	0.23 (0.25) ¹³	0.87 (0.95) ¹³	642	2.2	±0.009	45.45
[P _{6,6,6,14}][NTf ₂]	46.1	0.37	0.27	0.83	648	1.1	±0.005	90.91
[N _{1,8,8,8}][NTf ₂]	45.9	0.33	0.23	0.87	646	1.5	±0.008	66.67

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Analysis of isotope effects in NMR one-bond indirect nuclear spin–spin coupling constants in terms of localized molecular orbitals

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The numerical values given for the reduced coupling constants, K_{X-H} (in $10^{18} \text{ J}^{-1} \text{ T}^2$), and for the changes in the reduced coupling constants, ΔK_{X-H} (in $10^{18} \text{ J}^{-1} \text{ T}^2$), in Tables 2 to 7 and Figures 1 to 4 have to be multiplied by the constant factor π . These changes have no influence on the discussion or conclusions of the paper as all values are equally affected.